

Chemical Bond: Attraction between the protons of one atom and the electrons of the next atom that attaches the atoms together. It is formed by a sharing or transferring of electrons. All atoms have a complete outer shell and are stable after a bond forms.

- Has Potential energy
- When bond is formed, energy is released
- When a bond is broken, energy is absorbed.

Electronegativity: An atom's attraction for electrons in a chemical bond. Measured on an arbitrary scale from 0.0 - 4.0.

Octet Rule: Atoms tend to lose, gain, or share electrons to have a complete outer shell, just like a noble gas.

Lewis Dot Structures: Shows what happens to valence electrons.

- Each side can only have two electrons.
- dot structure can hold a max of 8 electrons.
- Place one dot on each side before doubling up.

Ionic Bonding: A metal transfers one or more electrons to a nonmetal to form ions.

- Attraction between a cation (positive) and an anion (negative).
- An electronegativity difference of the two ions is 2.0 or more.
- An ionic substance is hard, a good conductor of heat and electricity (can move in gas and liquid but can't move in solid form), high melting and boiling point (stronger attraction between charged ions), and dissolves in polar substances such as water.
- More ionic the greater the difference between the electronegativity of the atoms are.

Covalent Bonding: Formed when two atoms (both nonmetals) share electrons.

- Electronegativity difference is less than 2.0.
- Two main types of Covalent Bonding: Polar Covalent Bond & Nonpolar Covalent Bond.

Polar Covalent Bond: Electrons are shared between atoms of different elements (different electronegativity value), thus electrons are shared unequally as one element will exert a stronger pull on the other(s).

Nonpolar Covalent Bond: Electrons are shared between the atoms of the same element (same Electronegativity value) thus the electrons are shared equally.

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Minjie Lin



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