

Acids and Bases

Definitions	Acid	Base
1. Bronsted-Lowry	H ⁺ donor	H ⁺ acceptor
2. Lewis	e ⁻ acceptor	e ⁻ donor

Typical Acid/Base Reaction:



K_{eq} for this example K_a for this example:

$$K_{eq} = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{H}_2\text{O}][\text{HA}]} \quad K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$$

In general:

- Strong acid → Weak conj. base
- Weak acid → Strong conj. base
- Strong base → Weak conj. acid
- Weak base → Strong conj. acid

Predicting equilibria:

Start by labeling the pK_as in the reaction (see list)

Reaction Arrows:

- R₂N arrows point towards the side with the weaker acid (higher pK_a)

Is it reversible or irreversible?

- Look at the difference of pK_as between reactants/products
- $\Delta \text{pK}_a > 8 \Rightarrow$ irreversible rxn
- $\Delta \text{pK}_a < 8 \Rightarrow$ reversible rxn

Example: Predict equilibrium using pK_as, is it reversible or irreversible?



Acetylene **pK_a 25**

1. Label pK_as
2. Know arrows will point towards weaker acid (higher pK_a), so left
3. Is it reversible or irreversible? (25-16 > 8), so irreversible
4. Final arrows: \leftarrow

Things to know about acid/base reactions:

Equilibrium constant (K_{eq}) = ratio that tells the position of a chemical equilibrium

$$K_{eq} = \frac{[\text{products}]}{[\text{reactants}]}$$

Trends:

$K_{eq} > 1 \Rightarrow$ Favors products

$K_{eq} = 1 \Rightarrow$ At equilibrium

$K_{eq} < 1 \Rightarrow$ Favors reactants (non doesn't work)

Acid Dissociation Constant (K_a): measure of acid strength

$$K_a = K_{eq} [\text{H}_2\text{O}]$$

Trends:

$\uparrow K_a \Rightarrow$

- Strong acid
- Full dissociation
- High [products] (large numerator)

$\downarrow K_a \Rightarrow$

- Weak acid
- Partial/incomplete dissociation
- High [reactants] (large denominator)

$$\text{pK}_a = -\log K_a$$

Trends:

$\downarrow \text{pK}_a =$ Strong acid $\uparrow \text{pK}_a =$

Weak acid $\downarrow \text{pK}_a =$

- $\text{pK}_a < 1 \Rightarrow$ Strong acid
- $\text{pK}_a 1-3 \Rightarrow$ Moderately strong acid
- $\text{pK}_a 3-5 \Rightarrow$ Weak acid
- $\text{pK}_a > 5 \Rightarrow$ Very weak acid

pK_as to Know

- sp³ C-H 51
- sp² C-H 44
- sp C-H 25
- Water H₂O = 16 (15.7)
- Ethanol OH = 16 (15.9)
- Ammonia NH₃ 38
- Acetic Acid CO₂H = 5 (4.76)
- HCl -7

Study Guide Chemistry Acids And Bases

Clyde Metz



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Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement Organic Chemistry,

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