

1. Sulfur dioxide (SO₂)

- a. Phase: colorless
- b. Boiling point: 10°C
- c. Density: 2.93 g/L
- d. Solubility: 40 g/100 mL water at 20°C
- e. Chemical formula: SO₂
- f. Molecular weight: 64.06 g/mol

2. Write a balanced chemical equation for the reaction of SO₂ with H₂O to form H₂SO₃. (Sulfurous acid)3. Write a balanced chemical equation for the reaction of SO₂ with H₂O to form H₂SO₄.4. Write a balanced chemical equation for the reaction of SO₂ with H₂O to form H₂SO₄. (Sulfuric acid)5. Write a balanced chemical equation for the reaction of SO₂ with H₂O to form H₂SO₄.

Study Guide Chemistry Unit 8 Solutions

Robert J. Ouellette, J. David Rawn



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