

Ionic and Covalent Bonds Study Guide

Be able to define the following terms:

- | | | |
|-------------------|-----------------|--------------------|
| ➤ covalent bond | ➤ anion | ➤ polyatomic ion |
| ➤ Lewis structure | ➤ cation | ➤ hydrogen bonding |
| ➤ molecule | ➤ chemical bond | ➤ valence electron |
| ➤ polar bond | ➤ ionic bond | |
| ➤ non-polar bond | ➤ ion | |

1. A group of covalently bonded atoms that acts together as one charged atom is a polyatomic ion.
2. A chemical bond that occurs when atoms share electrons is a(n) covalent bond.
3. The octet rule tells you how many electrons an atom must gain, lose, or share to become stable.
4. The oxidation number of an atom is shown with a roman numeral in parenthesis.
5. How many electrons are needed in the outer energy levels of most atoms for the atom to be chemically stable? **8 (Eight)**
6. Why do the noble gases NOT form compounds readily? **They have 8 electrons in their outer electron shell.**
7. What is the total number of atoms in the compound $\text{Ca}(\text{ClO}_3)_2$? **9 (Nine)**
8. Explain why it is correct to say that a compound like calcium carbonate, CaCO_3 , contains both ionic and covalent bonds. **Carbonate is a polyatomic ion (covalent bonded) and then it forms an ionic bond with Calcium.**
9. Name or write the formulas for the following ionic compounds:
 - a. Al_2O_3 **Aluminum Oxide**
 - b. NH_4Cl **Ammonium Chloride**
 - c. FeO **Iron (II) Oxide**
 - d. manganese (V) oxide **Mn_2O_5**
 - e. calcium chloride **CaCl_2**
 - f. lead (IV) nitrate **$\text{Pb}(\text{NO}_3)_4$**

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