

Name _____ Date _____ Class _____

CHAPTER 4 STUDY GUIDE FOR CONTENT MASTERY**The Structure of the Atom****Section 4.1 Early Theories of Matter***In your textbook, read about the philosophers, John Dalton, and defining the atom.**For each statement below, write true or false.*

- false** _____ 1. Ancient philosophers regularly performed controlled experiments.
- true** _____ 2. Philosophers formulated explanations about the nature of matter based on their own experiences.
- true** _____ 3. Both Democritus and Dalton suggested that matter is made up of atoms.
- true** _____ 4. Dalton's atomic theory stated that atoms separate, combine, or rearrange in chemical reactions.
- false** _____ 5. Dalton's atomic theory stated that matter is mostly empty space.
- false** _____ 6. Dalton was correct in thinking that atoms could not be divided into smaller particles.
- true** _____ 7. Dalton's atomic theory stated that atoms of different elements combine in simple whole-number ratios to form compounds.
- true** _____ 8. Dalton thought that all atoms of a specific element have the same mass.
- false** _____ 9. Democritus proposed that atoms are held together by chemical bonds, but no one believed him.
- true** _____ 10. Dalton's atomic theory was based on careful measurements and extensive research.
- false** _____ 11. There are no instruments powerful enough to measure atoms so that they can be seen.
- true** _____ 12. The smallest particle of an element that retains the properties of that element is called an atom.

Name _____ Date _____ Class _____

CHAPTER 4 STUDY GUIDE FOR CONTENT MASTERY**Section 4.2 Subatomic Particles and the Nuclear Atom***In your textbook, read about discovering the electron and the nuclear atom.**For each item in Column A, write the letter of the matching item in Column B.*

Column A	Column B
5. _____ 1. Proposed the nuclear atomic model	a. Thomson
6. _____ 2. Determined the mass-to-charge ratio of an electron	b. Millikan
7. _____ 3. Calculated the mass of an electron	c. Rutherford

Draw and label a diagram of each atomic model.

4. **plum pudding model**

Drawings should look like a ball of chocolate chip cookie dough. The chocolate chips should be labeled with negative charge or as electrons. The dough should be labeled as evenly distributed positive charges.

5. **no-lose atomic model**

Drawings should look like a peach with a pit. The pit should be labeled nucleus and should include labeled protons and neutrons. The outer circle of the peach should be labeled electrons.

*In your textbook, read about the discovery of protons and neutrons.**Complete the following table of protons, electrons, and neutrons characteristics.*

Particle	Symbol	Location	Relative Charge	Relative Mass
1. Proton	p^+	In the nucleus	$1+$	1
2. Neutron	n^0	In the nucleus	0	1
3. Electron	e^-	In the space surrounding the nucleus	$1-$	$1/1840$

Study Guide Answers The Elements Chemistry

Jill Meryl Levy



Study Guide Answers The Elements Chemistry:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes. Also includes solutions and explanations to odd-numbered exercises.

AP Biology Study Guide Sundar Nathan, 2009-11 Sundar Nathan received a Bachelor's degree in Electrical Engineering from Anna University Chennai, India, and a Masters degree in Biomedical Engineering from the University of Texas at Austin. Working for over a year with a team of talented PhDs, MPhils, and MScs from all over the world, Sundar compiled this comprehensive study guide to help students prepare diligently, understand the concepts, and crush the AP Bio Test.

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane, CH_4 , from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18-chapter text designed to be used with Basics for Chemistry textbook. Each chapter contains Overview, Topical Outline, Skills, and Common Mistakes, which are all keyed to the textbook for easy cross-reference. The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms, a summary of general concepts, and a list of numerical exercises, while the Topical Outline provides the subtopic heads that carry the corresponding chapter

and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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