

\* Every 60 minutes, half of "A" converts to products.

8. For a reaction:  $aA \rightarrow \text{Products}$ ,  $[A]_0 = 6.0 \text{ M}$ , and the first two half-lives are 60. and 120. minutes, respectively. Calculate  $[A]$  at  $t = 99.9$  minutes. Ignore significant figures for this problem. (Hint: At time 60. minutes,  $[A] = 3.0 \text{ M}$  and at time 120 minutes,  $[A] = 1.5 \text{ M}$ . Use this to determine the order of the reaction.)

	$[A]$	Time	1st Order Run $K = -0.0116 \text{ 1/min}$
A) 5.4 M	6.0	0	$\ln[A] = -0.0116 \text{ 1/min} (99.9 \text{ min}) + \ln(6.0)$
B) 1.9 M	3.0	60	$\ln[A] = 0.633$
C) 1.0 M	1.5	120	$[A] = e^{0.633} \rightarrow [A] = 1.9 \text{ M}$
D) 0.65 M			
E) none of these			

Order  $r^2 = 0.964$  / 1st Order  $r^2 = 1.00$  / 2nd Order  $r^2 = 0.964$

Use the following to answer questions 9-13:

For the reaction  $2\text{N}_2\text{O}_5(\text{g}) \rightarrow 4\text{NO}_2(\text{g}) + \text{O}_2(\text{g})$ , the following data were collected:

$t$ (minutes)	$[\text{N}_2\text{O}_5]$ (mol/L)
0	$1.24 \times 10^{-2}$
10.	$0.92 \times 10^{-2}$
20.	$0.68 \times 10^{-2}$
30.	$0.50 \times 10^{-2}$
40.	$0.37 \times 10^{-2}$
50.	$0.28 \times 10^{-2}$
70.	$0.15 \times 10^{-2}$

Order  $r^2 = 0.910$   
 1st order  $r^2 = 0.999$   
 2nd order  $r^2 = 0.902$   
 Regression Information  
 Slope =  $K = -0.0301 \text{ 1/min}$

9. The order of this reaction in  $\text{N}_2\text{O}_5$  is

- A) 0  
 B) 1  
 C) 2  
 D) 3  
 E) none of these

10. The concentration of  $\text{O}_2$  (a product) at  $t = 10.$  minutes is (Hint: Look at the data table at 10. minutes. How much reactant has turned into product in the first 10 minute?)

- A)  $2.0 \times 10^{-4} \text{ mol/L}$   
 B)  $3.2 \times 10^{-3} \text{ mol/L}$   
 C)  $1.6 \times 10^{-3} \text{ mol/L}$   
 D)  $6.4 \times 10^{-3} \text{ mol/L}$   
 E) none of these

$$\begin{array}{r} 0.0124 \text{ 0 min} \\ - 0.0092 \text{ 10 min} \\ \hline 0.0032 \Delta[\text{N}_2\text{O}_5] \end{array} \quad \frac{0.0032 [\text{N}_2\text{O}_5]}{1} \times \frac{1 \text{ O}_2}{2 \text{ N}_2\text{O}_5} = 0.0016 [\text{O}_2]$$

11. The initial rate of decomposition of  $\text{N}_2\text{O}_5$  for this reaction is approximately ( $\text{Rate}_{\text{decomposition}} = k[\text{N}_2\text{O}_5]$ )

- A)  $6.4 \times 10^{-4} \text{ mol/L} \cdot \text{min}$   
 B)  $3.2 \times 10^{-4} \text{ mol/L} \cdot \text{min}$   
 C)  $1.24 \times 10^{-2} \text{ mol/L} \cdot \text{min}$   
 D)  $1.6 \times 10^{-4} \text{ mol/L} \cdot \text{min}$   
 E)  $2.6 \times 10^{-4} \text{ mol/L} \cdot \text{min}$

$$R_0 = \frac{\Delta[\text{N}_2\text{O}_5]}{\Delta \text{time}}$$

$$R_0 = \frac{0.0032 \text{ mol/L}}{10 \text{ min}}$$

$$R_0 = 0.00032 \text{ mol/L} \cdot \text{min}$$

# Nuclear Chemistry K Answers

**RJ Shavelson**



## **Nuclear Chemistry K Answers:**

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