

Chapter 19 – Nuclear Chemistry

This packet contains background and explanatory information on Nuclear Chemistry in addition to practice handouts. The information should be read and used to complete all of the practice included. I will be collecting this packet and giving points for the completion of the contents. You may want to consult your textbook (Chapter 19) or online resources for addition explanations and examples if needed.

Chapter Objectives: The student will (be able to)...

- Compare and contrast **alpha**, beta, and gamma radiation.
- Identify alpha decay, beta decay, fission, and fusion reactions.
- Balance nuclear equations.
- Solve half-life calculations.

Part I: Radioactivity - What is it? <http://www.darvill.clara.net/nucrad/index.htm>

All substance are made of **atoms**. These have **electrons** (e) around the outside, and a **nucleus** in the middle. The nucleus consists of **protons** (p) and **neutrons** (n), and is extremely small. (Atoms are almost entirely made of empty space!)

In some types of atom, the **nucleus is unstable**, and will **decay** into a more stable atom. This **radioactive decay** is completely **spontaneous**. The energy that is released from the nucleus of the atom is **radiation**.

When an unstable nucleus **decays**, there are **three** ways that it can do so. It may give out:-

- an **alpha particle** (we use the symbol α)
- a **beta particle** (symbol β)
- a **gamma ray** (symbol γ)

Many radioactive substances **emit** α particles and β particles as well as γ rays. In fact, you won't find a pure γ source: anything that gives off γ rays will also give off α and/or β too.

A. Alpha Particles – More Information

- Alpha particles are made of **2 protons and 2 neutrons**.
- This means that they have a **charge of +2**, and a **mass of 4** (the mass is measured in "atomic mass units", where each proton & neutron = 1)

We can write them as ${}^4_2\alpha$, or, because they're the same

as a helium nucleus, ${}^4_2\text{He}$.

- Alpha particles are relatively **slow and heavy**.
- They have a **low penetrating power** - you can stop them with just a **sheet of paper**.
- Because they have a large charge, alpha particles **ionize** (pull electrons from) other atoms strongly.



Alpha particle:
2 protons
& 2 neutrons

Nuclear Chemistry Packet Answer Key

Puja Mehta



Nuclear Chemistry Packet Answer Key:

Essays in Contemporary Chemistry Gerhard Quinkert, 2001 In this volume internationally renowned chemists recount their roles in the progress of chemistry research toward elucidation of biological processes Beginning with a historical perspective on the development of X ray crystallography the reader is regaled with first hand accounts of research milestones Included are descriptions of the cutting edge nuclear magnetic resonance and electron paramagnetic resonance spectroscopic techniques the dynamics of ultra fast reactions and the central problem of molecular oxygen activation in biological processes The roles of chiral auxiliaries in organic synthesis and of photochemistry in natural product synthesis are told and innovations in combinatorial synthesis approaches are described Contemporary issues in molecular recognition and modulation of molecular function are addressed concluding with a missive regarding how the frontiers of medical ethics may be breached by molecular manipulations The contributors who number among the finest scientists in the world including two Nobel Prize winners are Peter B Dervan Jack D Dunitz Christian Griesinger Jean Marie Lehn Thomas F Prisner Gerhard Quinkert Peter G Schultz Helmut Schwarz Dieter Seebach and Ahmed Zewail Additionally there is a prologue by Albert Eschenmoser for whom this collection was conceived and an epilogue that contains facsimiles of notes from his landmark lecture Synthesis of Co Enzyme B12 A Vehicle for Teaching Organic Synthesis This book is definitely a must for all who want to read or to read again where we stand in our chemical comprehension of the fascinating relationship between chemical structure and biological processes how we got here and what the future might hold

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Quantum Chemistry and Dynamics of Excited States Leticia González, Roland Lindh, 2021-02-01 An introduction to the rapidly evolving methodology of electronic excited states For academic researchers postdocs graduate and undergraduate students Quantum Chemistry and Dynamics of Excited States Methods and Applications reports the most updated and accurate theoretical techniques to treat electronic excited states From methods to deal with stationary calculations through time dependent simulations of molecular systems this book serves as a guide for beginners in the field and knowledge seekers alike Taking into account the most recent theory developments and representative applications it also covers the often overlooked gap between theoretical and computational chemistry An excellent reference for both

researchers and students **Excited States** provides essential knowledge on quantum chemistry an in depth overview of the latest developments and theoretical techniques around the properties and nonadiabatic dynamics of chemical systems Readers will learn Essential theoretical techniques to describe the properties and dynamics of chemical systems Electronic Structure methods for stationary calculations Methods for electronic excited states from both a quantum chemical and time dependent point of view A breakdown of the most recent developments in the past 30 years For those searching for a better understanding of excited states as they relate to chemistry biochemistry industrial chemistry and beyond **Quantum Chemistry and Dynamics of Excited States** provides a solid education in the necessary foundations and important theories of excited states in photochemistry and ultrafast phenomena Christian Home Educators' Curriculum Manual Cathy Duffy, 2000 Reviews goal setting what to teach learning styles how to teach planning and record keeping resource addresses Cover

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