

Nuclear Chemistry Worksheet

Identify the following types of nuclear radiation

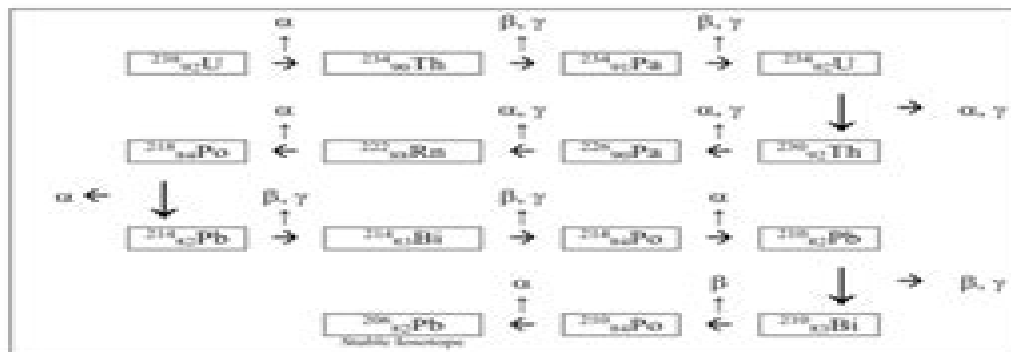
- $\frac{0}{-1}e$ _____
- $\frac{4}{2}He$ _____
- $\frac{0}{0}\gamma$ _____
- Has no mass or charge, almost always follows α or β decay _____
- Occurs when a neutron breaks apart and forms a proton _____
- Least penetrating nuclear decay _____
- Most damaging nuclear decay to the human body _____
- Nuclear decay that can be stopped by skin or paper. _____
- Nuclear decay that can be stopped by aluminum. _____
- As a result, a proton is transformed into a neutron _____
- The opposite of beta decay _____
- Type of decay isotopes $Z > 83$ will undergo _____

Complete the following nuclear equations.

13. ${}_{19}^{42}\text{K} \rightarrow {}_{-1}^0\text{e} + \underline{\hspace{2cm}}$
14. ${}_{4}^9\text{Be} \rightarrow {}_{4}^9\text{Be} + \underline{\hspace{2cm}}$
15. ${}_{3}^6\text{Li} \rightarrow {}_{2}^4\text{He} + \underline{\hspace{2cm}}$
16. ${}_{94}^{239}\text{Pu} \rightarrow {}_{2}^4\text{He} + \underline{\hspace{2cm}}$
17. ${}_{92}^{235}\text{U} \rightarrow \underline{\hspace{2cm}} + {}_{90}^{231}\text{Th}$
18. $\underline{\hspace{2cm}} \rightarrow {}_{56}^{142}\text{Ba} + {}_{36}^{91}\text{Kr} + 3 {}_{0}^1\text{n}$

Nuclear Decay Series

The figure maps the radioactive decay of uranium-238 to lead-206. Use the figure to answer the following questions.



19. How many alpha particles are produced as one atom of uranium-238 decays to an atom of lead-206?

20. How many beta particles? _____
21. Explain why lead-206 is a stable isotope. _____

Nuclear Chemistry Worksheet Chapter 21

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



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Solutions Manual for Principles of Physical Chemistry, 3rd Edition Hans Kuhn, David H. Waldeck, Horst-Dieter Försterling, 2024-10-29 This is a Solutions Manual to Accompany with solutions to the exercises in the main volume of Principles of Physical Chemistry Third Edition This book provides a unique approach to introduce undergraduate students to the concepts and methods of physical chemistry which are the foundational principles of Chemistry The book introduces the student to the principles underlying the essential sub fields of quantum mechanics atomic and molecular structure atomic and molecular spectroscopy statistical thermodynamics classical thermodynamics solutions and equilibria electrochemistry kinetics and reaction dynamics macromolecules and organized molecular assemblies Importantly the book develops and applies these principles to supramolecular assemblies and supramolecular machines with many examples from biology and nanoscience In this way the book helps the student to see the frontier of modern physical chemistry developments The book begins with a discussion of wave particle duality and proceeds systematically to more complex chemical systems in order to relate the story of physical chemistry in an intellectually coherent manner The topics are organized to correspond with those typically given in each of a two course semester sequence The first 13 chapters present quantum mechanics and spectroscopy to describe and predict the structure of matter atoms molecules and solids Chapters 14 to 29 present statistical thermodynamics and kinetics and applies their principles to understanding equilibria chemical transformations macromolecular properties and supramolecular machines Each chapter of the book begins with a simplified view of a topic and evolves to more rigorous description in order to provide the student and instructor flexibility to choose the level of rigor and detail that suits them best The textbook treats important new directions in physical chemistry research including chapters on macromolecules principles of interfaces and films for organizing matter and supramolecular machines as well as including discussions of modern nanoscience spectroscopy and reaction dynamics throughout the text Solutions Manual for Principles of Physical Chemistry, 3rd Edition, Solutions Manual Hans Kuhn, David H. Waldeck, Horst-Dieter Försterling, 2024-10-25 This is a Solutions Manual to Accompany with solutions to the exercises in the main volume of Principles of Physical Chemistry Third Edition This book provides a unique approach to introduce undergraduate students to

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