





- Your work is to be done on separate sheets of paper.
- · All of the papers should be stapled together in the upper left hand corner with this question sheet at the back of the packet.
- Each of your responses should be numbered the same way as it appears on this packet.

Conceptual Ouestions

- State the law of conservation of mass.
- 2) Why are atoms unstable?
- 3) How do unstable atoms gain stability?
- 4) A radioactive element having atomic number 82 and atomic mass 214 loses a beta particle, 0,-1β. What is the resulting particle?
- 5) What type of decay is the reaction below, and what isotope completes this nuclear equation? $27.13AI + 2.1H \rightarrow 4.2He + 7$

Practice Questions

6) Complete this chart on the three types of radioactive decay

Type	Symbol	Mass (amu)	Charge	
Alpha	G 22			
Beta				
Gamma				

7) If the radioactive atom 238, 92U emits an alpha particle, the atom remaining is represented by

(A) 238, 9211 (B) 236, 92Th (C) 234, 92U (D) 234, 99Th (E) 234, 91Pa

What term completes this equation? ^{30,14}Si + ? → 32,16S + ^{0,-1}e

 $(D)^{-q_{i-1}}e^{-t}$ (A)4.2He (C)*-!H (B)2.4H

9) What particle is emitted when lead-214, 214, 82Pb, undergoes natural radioactive decay to bismuth-214, 214, 83Bi?

(A) an alpha particle, 42He a proton, 13H. (C)

(B) a beta particle, %-1e (D) a neutron, 1,0n

Extension Question

- 10) Recently Alexander Valterovich Litvinenko was poisoned by radioactive polonium-210.
 - Research who he was and why he was poisoned.
 - Find out what polonium-210 does to people.

Nuclear Chemistry Worksheet Chapter 25

Tobias Bleicker

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