

**Chapter 10 Nuclear Chemistry**

7. How is the product isotope different from the reactant isotope in alpha decay? The product isotope has two fewer protons and two fewer neutrons than the reactant isotope.
8. Circle the letters that identify each event that takes place during beta decay.
- A proton decomposes into a neutron and an electron.
  - A neutron decomposes into a proton and an electron.
  - An electron is emitted from the nucleus.
  - A neutron is emitted from the nucleus.
9. Why are beta particles more penetrating than alpha particles?  
Beta particles have a smaller mass and a faster speed.
10. Is the following sentence true or false? All nuclear radiation consists of charged particles. false
11. What is a gamma ray? A gamma ray is a penetrating ray of energy emitted by an unstable nucleus.
12. How fast do gamma rays travel through space?  
Gamma rays travel through space at the speed of light.
13. Complete the following table about nuclear radiation.

Characteristics of Nuclear Radiation			
Radiation Type	Charge	Mass (amu)	Usually Stopped By
Alpha particle	-1	4	Paper or clothing
Beta particle	-1	$\frac{1}{1836}$	Aluminum sheet
Gamma ray	0	0	Several meters of concrete

**Effects of Nuclear Radiation (pages 296-297)**

14. How does nuclear radiation affect atoms? Nuclear radiation can ionize atoms.
15. Is the following sentence true or false? One potential danger of radon gas is that prolonged exposure to it can lead to lung cancer.  
true

**Detecting Nuclear Radiation (page 297)**

16. Name two devices that are used to detect nuclear radiation.
- a. Geiger counters      b. Film badges

## Nuclear Energy Answer Key Section 10 3

**D Kirk**

### **Nuclear Energy Answer Key Section 10 3:**

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*Energy Abstracts for Policy Analysis*, 1978

*Energy Science* John Andrews, John G. Andrews, Nicholas Alfred Jelley, Nick Jelley, 2022 Can climate change be tackled with the current energy technologies and policies available? Are renewable energy sources now economically competitive with fossil fuels? How much power can be provided by all the various forms of low carbon energy? Energy Science introduces the latest energy technologies, explains the physical principles underlying each technology and discusses their environmental, economic and social impacts. Covering fossil fuels, renewables and nuclear energy, this book provides you with tools to evaluate the key sources of energy available and introduces potential solutions to the energy problems facing us today. New to this edition: Greater emphasis on the challenges and urgent actions required in the transition from fossil fuels to low carbon sources of energy; explanations of recent developments in solar, PV, off shore wind, deploying renewables and battery technology; More international case studies looking at real world energy issues and applications of energy technology. Increased focus on making the text engaging for a wide range of students with flexible treatment of the more mathematical content.

**Book jacket**

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**Principles of Quantum Chemistry** David V. George, 2013-10-22 Principles of Quantum Chemistry focuses on the application of quantum mechanics in physical models.

and experiments of chemical systems This book describes chemical bonding and its two specific problems bonding in complexes and in conjugated organic molecules The very basic theory of spectroscopy is also considered Other topics include the early development of quantum theory particle in a box general formulation of the theory of quantum mechanics and treatment of angular momentum in quantum mechanics The examples of solutions of Schroedinger equations approximation methods in quantum chemistry symmetry in chemistry and molecular orbital theory are also covered This publication is recommended for students taking undergraduate and graduate courses in quantum chemistry

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