

# Study Guide Answers

## Chapter 21

Name: \_\_\_\_\_

Date: \_\_\_\_\_ Per: \_\_\_\_\_

- The splitting of the nucleus into lighter nuclei is called fission.
- When light mass nuclei combine to form a heavier, more stable nucleus this is called fusion.
- Write an equation for the alpha decay of bismuth-213.  $^{213}_{83}\text{Bi} \rightarrow ^4_2\text{He} + ^{209}_{81}\text{Tl}$
- Write an equation for the beta decay of silver-106.  $^{106}_{47}\text{Ag} \rightarrow ^0_{-1}\text{e} + ^{106}_{48}\text{Cd}$
- Describe the difference between beta decay and electron capture. In beta decay, an electron is emitted from the nucleus as a neutrons transforms into a proton. In electron capture (or K capture or beta capture) an electron is taken in from the electron cloud by the nucleus to convert a proton into a neutron.

- The mass of a lithium-7 atom\* is 7.01600amu. Calculate it mass defect.

$$p^+ \text{ mass} = 3 \times 1.007276 \text{ amu} = 3.021828 \text{ amu}$$

$$n^0 \text{ mass} = 4 \times 1.008665 \text{ amu} = 4.034660 \text{ amu}$$

$$e^- \text{ mass} = 7 \times 0.0005486 \text{ amu} = 0.0038402 \text{ amu} \text{ (* mass is for atom, not nucleus)}$$

$$\text{Theoretical Mass} = 7.0581338 \text{ amu}$$

$$\text{Mass defect} = 7.0581338 \text{ amu} - 7.01600 \text{ amu} = \boxed{0.04213 \text{ amu}}$$

Proton Mass 1.007276amu  
Neutron Mass 1.008665amu  
Electron Mass 0.0005486amu

- Calculate the nuclear binding energy of one lithium-6 atom. The actual measured atomic mass of lithium-6 is 6.015amu.

$$p^+ \text{ mass} = 3 \times 1.007276 \text{ amu} = 3.021828 \text{ amu}$$

$$n^0 \text{ mass} = 3 \times 1.008665 \text{ amu} = 3.025995 \text{ amu}$$

$$e^- \text{ mass} = 7 \times 0.0005486 \text{ amu} = 0.0038402 \text{ amu} \text{ (* mass is for atom, not nucleus)}$$

$$\text{Theoretical Mass} = 6.0496838 \text{ amu}$$

$$\text{Mass defect} = 6.0496838 \text{ amu} - 6.015 \text{ amu} = 0.0346838 \text{ amu} \text{ (must be expressed in kg to use } E=mc^2\text{)}$$

$$\frac{0.0346838 \text{ amu}}{6.022 \times 10^{23} \text{ amu}} \times \frac{1 \text{ gram}}{1000 \text{ grams}} = 5.723812687 \times 10^{-29} \text{ kg}$$

$$E = mc^2 = 5.723812687 \times 10^{-29} \text{ kg} \cdot (3.00 \times 10^8 \text{ m/s})^2 = \boxed{5.15 \times 10^{-13} \text{ J}}$$

*Since I used 3 sig figs for "c", I rounded the final answer to 3 sig figs.*

- Oxygen has an unstable isotope O-17 that has a mass of 17.00454. Calculate the binding energy of the oxygen nucleus in MeV.

$$p^+ \text{ mass} = 8 \times 1.007276 \text{ amu} = 8.058208 \text{ amu}$$

$$n^0 \text{ mass} = 9 \times 1.008665 \text{ amu} = 9.077985 \text{ amu}$$

$$e^- \text{ mass} = 8 \times 0.0005486 \text{ amu} = 0.0043888 \text{ amu} \text{ (* mass is for atom, not nucleus)}$$

$$\text{Theoretical Mass} = 17.1405818 \text{ amu}$$

$$\text{Mass defect} = 17.1405818 \text{ amu} - 17.00454 = 0.1360418 \text{ amu} \text{ (must be expressed in kg to use } E=mc^2\text{)}$$

$$\frac{0.1360418 \text{ amu}}{6.022 \times 10^{23} \text{ amu}} \times \frac{1 \text{ gram}}{1000 \text{ grams}} = 2.25908004 \times 10^{-28} \text{ kg}$$

$$E = mc^2 = 2.25908004 \times 10^{-28} \text{ kg} \cdot (3.00 \times 10^8 \text{ m/s})^2 = 2.03382036 \times 10^{-11} \text{ J}$$

$$1 \text{ eV} = 1.60 \times 10^{-19} \text{ joules (I'd give this value to you in the test)}$$

$$\frac{2.03382036 \times 10^{-11} \text{ J}}{1.60 \times 10^{-19} \text{ J}} \times \frac{1 \text{ MeV}}{1000000 \text{ eV}} = \boxed{127 \text{ MeV}}$$

*Since I used 3 sig figs for "c" and the eV conversion, I rounded the final answer to 3 sig figs.*

# Nuclear Chemistry Packet Answers Chapter 21

**O García**



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