#### Predicting Precipitate Formation - Additional Practice

- Using the solubility rules, determine whether each of the following compounds is soluble in water:
  - a. Calcium sulfate
  - b. Hydrogen sulfide
  - c. Sodium nitrate
  - d. Copper (II) chloride
  - e. Magnesium carbonate
  - Sodium hydroxide
- Complete each of the following word equations. Using the solubility rules, determine whether a precipitate forms. If a precipitate forms, write a balanced chemical equation.
  - a. Magnesium nitrate and lithium sulfate
  - b. Lead (II) acetate and silver nitrate
  - Sodium bromide and sodium nitrate.
  - d. Silver nitrate and ammonium phosphate
  - e. Magnesium hydroxide and calcium nitrate
  - f. Potassium carbonate and barium chloride
  - g. Silver nitrate and potassium chloride
  - h. Lead (II) nitrate and potassium sulfide

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Elena Hilario, John F. MacKay

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