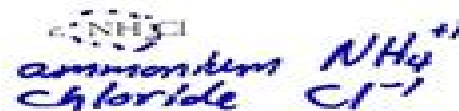
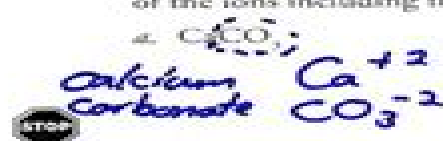


15. Bromine forms polyatomic ions with structures similar to those of chlorine. Using the chlorine family of polyatomic ions as a model, predict the name of the BrO_4^- ion.

ClO_4^- is perchlorate, therefore, BrO_4^- is perbromate.

16. Identify the polyatomic ion in each of these ionic compounds. Write out the name and formula of the ions including their charges.



Model 3 – Ternary Ionic Compounds

Compound Name	Ion Symbols and Charges		Chemical Formula
Ammonium phosphate	NH_4^{+1}	PO_4^{3-}	$(\text{NH}_4)_3\text{PO}_4$
Barium nitrite	Ba^{2+}	NO_2^{-1}	$\text{Ba}(\text{NO}_2)_2$
Ammonium sulfate	NH_4^{+1}	SO_4^{2-}	$(\text{NH}_4)_2\text{SO}_4$
Aluminum carbonate	Al^{3+}	CO_3^{2-}	$\text{Al}_2(\text{CO}_3)_3$
Iron(III) hydroxide	Fe^{3+}	OH^{-1}	$\text{Fe}(\text{OH})_3$
Potassium nitrate	K^{+}	NO_3^{-1}	KNO_3

17. How are ternary ionic compounds in Model 3 different from binary ionic compounds (NaCl , MgO , CaBr_2 , etc.) that you've seen previously? *Hint:* Consider the meaning of the word "binary."

Ternary ionic compounds contain more than two elements.
 Binary ionic compounds contain only two elements.

18. Consider the compound iron(III) hydroxide in Model 3.

- a. How many hydroxide ions (OH^{-1}) are combined with an iron(III) ion (Fe^{3+})?

There are three hydroxide (OH^{-1}) for every one iron(III) (Fe^{3+}) ion.

- b. Is your answer to part a the only combination of iron(III) and hydroxide that should exist in nature? Explain.

Yes. The ratio of ~~three~~ ^{one} Fe^{3+} and three OH^{-1} is the only ratio that can balance the charge of the two ions.

19. Consider the compound barium nitrite in Model 3.

- a. What does the subscripted "2" inside the parentheses of the chemical formula tell you about the compound?

The "2" inside of the parentheses indicates that there are two oxygen atoms in the nitrite (NO_2^{-1}) polyatomic ion.

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**Alper Baba, Jochen Bundschuh, D.
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