

Chapter 19 – Nuclear Chemistry

This packet contains background and explanatory information on Nuclear Chemistry in addition to practice handouts. The information should be read and used to complete all of the practice included. I will be collecting this packet and giving points for the completion of the contents. You may want to consult your textbook (Chapter 19) or online resources for addition explanations and examples if needed.

**Chapter Objectives:** The student will (be able to)...

- Compare and contrast **alpha**, beta, and gamma radiation.
- Identify alpha decay, beta decay, fission, and fusion reactions.
- Balance nuclear equations.
- Solve half-life calculations.

### Part I: Radioactivity - What is it? <http://www.darvill.clara.net/nucrad/index.htm>

All substance are made of **atoms**. These have **electrons** (e) around the outside, and a **nucleus** in the middle. The nucleus consists of **protons** (p) and **neutrons** (n), and is extremely small. (Atoms are almost entirely made of empty space!)

In some types of atom, the **nucleus** is **unstable**, and will **decay** into a more stable atom. This **radioactive decay** is completely **spontaneous**. The energy that is released from the nucleus of the atom is **radiation**.

When an unstable nucleus **decays**, there are **three** ways that it can do so. It may give out:-

- an **alpha particle** (we use the symbol  $\alpha$ )
- a **beta particle** (symbol  $\beta$ )
- a **gamma ray** (symbol  $\gamma$ )

Many radioactive substances **emit**  $\alpha$  particles and  $\beta$  particles as well as  $\gamma$  rays. In fact, you won't find a pure  $\gamma$  source: anything that gives off  $\gamma$  rays will also give off  $\alpha$  and/or  $\beta$  too.

#### A. Alpha Particles – More Information

- Alpha particles are made of **2 protons** and **2 neutrons**.
- This means that they have a **charge** of **+2**, and a **mass** of **4** (the mass is measured in "atomic mass units", where each proton & neutron = 1)

We can write them as  ${}^4_2\alpha$ , or, because they're the same

as a helium nucleus,  ${}^4_2\text{He}$ .

- Alpha particles are relatively **slow** and **heavy**.
- They have a **low penetrating power** - you can stop them with just a **sheet of paper**.
- Because they have a large charge, alpha particles **ionize** (pull electrons from) other atoms strongly.



**Alpha particle:**  
**2 protons**  
**& 2 neutrons**

# Nuclear Chemistry Packet Answers

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novel developments to meet the current challenges in the field Further this monograph responds to a need for a systematic comparative treatise on nonadiabatic theories of quantum molecular dynamics which are of considerably higher complexity than the more traditional adiabatic approaches and are steadily gaining in importance This volume addresses a broad readership ranging from physics or chemistry graduate students to specialists in the field of theoretical quantum dynamics

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