

Unit 7: Reduction, Oxidation and ElectrochemistryChapter 17: Electrochemistry4.9: Oxidation-Reduction Reactions

Reduction-Oxidation Reactions (Redox Rxn): - chemical reactions where there is a transfer of electron(s).

Oxidation States (Oxidation Number): - a number that is arbitrary assigned to an atom in an element, molecule, and polyatomic ions to account for the number of electrons in *redox* reaction.

Rules for Assigning Oxidation Numbers:

1. Natural Elements (include diatomic and polyatomics) have an Oxidation Number of 0.

Examples: Na_(s), O_{2(g)}, O_{3(g)}, H_{2(g)}, F_{2(g)}, P_{4(s)}, and Hg_(l) all have an oxidation number of 0.

2. Single Atomic Ions have an Oxidation Number Equals to its Charge.

Example: K⁺ has an oxidation number of +1.

3. Oxygen in Binary Compound and Polyatomic Ions has an Oxidation Number of -2.

Examples: CO₂, CO, SO₃, SO₂, CO₃²⁻, SO₄²⁻ and H₂O all have an oxidation number of -2 for oxygen.

An exception occurs in peroxides (compound containing O₂²⁻) where the oxygen has an oxidation number of -1.

4. Hydrogen in Binary Compounds and Polyatomic Ions has an Oxidation Number of +1.

Examples: H₂O, HCl, HBr, CH₄, NH₄⁺ and H₂S all has an oxidation number of + 1 for hydrogen.

5. In Binary Compounds, the Atom that has the most attraction to electrons (HIGHER Electronegativity) is assigned the NEGATIVE Oxidation Number. This Negative Oxidation Number is the SAME as its Ionic Charge.

Examples: HF: F is more electronegative. Thus, F has an oxidation number of -1.

NH₃: N is more electronegative. Thus, N has an oxidation number of -3.

H₂S: S is more electronegative. Thus, S has an oxidation number of -2.

6. The Sum of all Oxidation Numbers is 0 for all Electrically Neutral Compounds.

Examples: CO: Oxidation numbers: O = -2; C = +2 [(+2) + (-2) = 0]

SO₂: Oxidation numbers: O = -2; S = +4 [(+4) + 2(-2) = 0]

7. The Sum of all Oxidation Numbers is Equal to the Overall Charge for all Polyatomic Ions.

Examples: CO₃²⁻: Oxidation numbers: O = -2; C = +4 [(+4) + 3(-2) = -2]

SO₄²⁻: Oxidation numbers: O = -2; S = +6 [(+6) + 4(-2) = -2]

Note: Same Atom from different chemical species might NOT have the Same Oxidation Number.

We write the charges of ionic species as *n*⁺ or *n*⁻. Oxidation numbers are written as +*n* or -*n*.

Oxidation Reduction Study Guide Answers

Xiaolong Qi



Oxidation Reduction Study Guide Answers:

This book delves into Oxidation Reduction Study Guide Answers. Oxidation Reduction Study Guide Answers is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Oxidation Reduction Study Guide Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Oxidation Reduction Study Guide Answers
- Chapter 2: Essential Elements of Oxidation Reduction Study Guide Answers
- Chapter 3: Oxidation Reduction Study Guide Answers in Everyday Life
- Chapter 4: Oxidation Reduction Study Guide Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Oxidation Reduction Study Guide Answers. The first chapter will explore what Oxidation Reduction Study Guide Answers is, why Oxidation Reduction Study Guide Answers is vital, and how to effectively learn about Oxidation Reduction Study Guide Answers.

3. In chapter 2, the author will delve into the foundational concepts of Oxidation Reduction Study Guide Answers. The second chapter will elucidate the essential principles that must be understood to grasp Oxidation Reduction Study Guide Answers in its entirety.

4. In chapter 3, this book will examine the practical applications of Oxidation Reduction Study Guide Answers in daily life. The third chapter will showcase real-world examples of how Oxidation Reduction Study Guide Answers can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Oxidation Reduction Study Guide Answers in specific contexts. This chapter will explore how Oxidation Reduction Study Guide Answers is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Oxidation Reduction Study Guide Answers. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Oxidation Reduction Study Guide Answers.

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