

Compounds

Word	Definition
Binary compound	A compound that consists of two elements.
Dipole attractions	The attraction of the δ^- end of one polar molecule to the δ^+ end of another polar molecule.
Dipole moment	An arrow along the line of symmetry in a polar molecule that shows the net direction that electrons are being pulled towards the partially negative end of the molecule.
Electronegativity difference	The difference in electronegativity between two elements in a bond.
Electrical conductivity (metals)	The ability of a substance to allow electrons to pass from atom to atom through itself from a source or electricity to an electrical ground.
Electrolyte	A solution containing dissolved ions that can conduct electricity through it.
Empirical formula	The simplest whole-number mole ratio of elements in a compound, used to write the formulas of ionic compounds.
Formula mass	The sum of the atomic masses of an element or compound, measured in grams/mole. Reported to the nearest tenth of a gram per mole.
Hydrogen bonds	The strong attraction of the H (δ^+) end of one polar molecule to the N, O or F (δ^-) end of another polar molecule. The two ends form temporary covalent bonds.
Intermolecular Attractive Force	The forces that hold molecules together in the solid and liquid phase. These are the forces that must be overcome to melt or boil a substance. Also called "van der Waal's forces".
Ionic compound	Compounds consisting of a metal and a nonmetal that are ionically bonded in a whole-number ratio.
London dispersion force	The weak attractive force caused by temporary dipoles in nonpolar molecules.
Metallic bond	A bond formed between metal atoms as they collectively share their conducting electrons evenly between metal kernels.
Molecular formula	The actual number of nonmetal atoms in a molecule, a whole-number multiple of the empirical formula.
Molecule	A particle made of nonmetal atoms covalently bonded together to form a distinct particle.
Network solid	A crystal lattice formed from covalently bonded nonmetal atoms with no distinct molecules.
Nonpolar molecule	A molecule with symmetrical electron distribution resulting in any polar bonds canceling each other out to yield no partially charged ends.
Percent composition	The formula mass of an element divided by the formula mass of the compound that contains the element multiplied by 100.
Polar molecule	A molecule with asymmetrical electron distribution resulting in partially charged ends.
Polyatomic ion	An ion formed by atoms bonding together in such a way that a net charge is formed.
Ternary compound	A compound that consists of three or more elements, usually containing a polyatomic ion.

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Marion E. Potter



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