

The following is a summary of most of the concepts, relationships, topics that you should know for this course. Note that this may include typographical errors because it has not been proof read and is a first draft, so please check all expressions with those in the book. I would only use this as a study guide and not as a replacement for studying the topic in depth within the book. Also note that I provide information about certain things I recommend you know how to do, calculate, derive, etc. *Prof. Musgrave*

Section 1.2a – Perfect gas law

Calculating relationships between p , V and T at one state and another.

The molecules in perfect gases experience no interactions.

Do not apply the perfect gas law to systems where it does not apply (liquids, solids, dense gases)

Section 1.2b – Kinetic model of gases

Calculating average speeds of gas molecules as a function of M and T .

Section 1.2c – Mixtures of gases

Calculating partial pressures from total pressure and composition (mole fractions)

Dalton's Law: The sum of the pressures of gases of each component alone in a vessel sum to the total pressure of the mixture of those gases.

Section 1.3a – The compression factor

The volumes of real gases are different from perfect gases by the compression factor.

The compression factor is related to the Virial equation of state

Section 1.4a – van der Waals equation

Corrections for the size of molecules and their interactions can be included to form a more accurate description of the equation of state $p(V,T)$

There is no closed form expression of $V(p,T)$ – solve for V numerically.

There are closed forms of $p(T,V)$ and $T(p,V)$ so inverting partial derivatives is sometimes convenient to determine derivatives of V with respect to T or p .

Section 1.4b – van der Waals equation features

The vdW equation of state differs from the ideal gas law at high p and low T (where the assumption of ideal behavior breaks down).

Chapter 1: Key Equations:

Perfect gas law: $pV = nRT$

Temperature scales: $T(K) = T(C) + 273.15$

Partial pressures: $p = \sum p_i$ and $p_i = x_i p = \frac{n_i}{n} p$

Van der Waals Equation of state: $p = nRT / (V - nb) - an^2 / V^2$

Section 2.1 – Work, heat and energy

The energy of a system is its capacity to do work

Work is the transfer of energy by moving against a force – it is the transfer of energy by organized motion.

Heat is the transfer of random, thermal motion. It transfers disordered energy.

An adiabatic process transfers no heat and thus no random energy or entropy.

An isothermal process involves no change in T .


Section 2.2 – Internal energy U

Internal energy (U) is the total energy of the system. It includes kinetic energy and potential energy and is a state of the system. It increases because either work is done on the system or heat is transferred to the system.

Section 2.2a – The internal energy of a perfect gas depends on T only – $U(T)$.

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**Alexander V. Vakhrushev, Reza
Haghi, J.V. de Julián-Ortiz**



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A Life Scientist's Guide to Physical Chemistry Marc R. Roussel, 2012-04-05 Demonstrates how the tools of physical chemistry can be applied to biological questions with numerous exercises and clearly worked examples

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A Laboratory Manual of Elementary Physical Chemistry Edward Mack, Wesley George France, 1928

Thermodynamics Problem Solving in Physical Chemistry Kathleen E. Murphy, 2020-03-23

Thermodynamics Problem Solving in Physical Chemistry Study Guide and Map is an innovative and unique workbook that guides physical chemistry students through the decision making process to assess a problem situation create appropriate solutions and gain confidence through practice solving physical chemistry problems The workbook includes six major sections with 20 30 solved problems in each section that span from easy single objective questions to difficult multistep analysis problems Each section of the workbook contains key points that highlight major features of the topic to remind students of what they need to apply to solve problems in the topic area Key Features Provides instructor access to a visual map depicting how all equations used in thermodynamics are connected and how they are derived from the three major energy laws Acts as a guide in deriving the correct solution to a problem Illustrates the questions students should ask themselves about the critical features of the concepts to solve problems in physical chemistry Can be used as a stand alone product for review of Thermodynamics questions for major tests

Physical Chemistry for Chemists and Chemical Engineers Alexander V. Vakhrushev, Reza Haghi, J.V. de Julián-Ortiz, 2018-09-03

This volume is based on different aspects of chemical technology that are associated with research and the development of theories for chemical engineers helping to bridge the gap between classical analysis and modern real life applications Taking an interdisciplinary approach the authors present the current state of the art technology in key materials with an emphasis on the rapidly growing technologies

Advances in Teaching Physical Chemistry Mark David Ellison, 2008

This book brings together the latest perspectives and ideas on teaching modern physical chemistry It includes perspectives from experienced and well known physical chemists a thorough review of the education literature pertaining to physical chemistry a thorough review of advances in undergraduate laboratory experiments from the past decade in depth descriptions of using computers to aid student learning and innovative ideas for teaching the fundamentals of physical chemistry This book will provide valuable insight and information to all teachers of physical chemistry

A Non-Traditional Guide to Physical Chemistry Robert Schiller, 2022-08-17

This book introduces in a non traditional way the laws of physical chemistry and its history starting in the 16th century It reveals to the reader how physical chemists try to understand chemical processes in terms of physical laws Hydrogen is the main focus of the book as its simplicity makes the relevant laws of nature easy to explain and its role in energetics in the near future is clear With the basics at hand the importance of hydrogen as a raw material in the industry and as an energy carrier in the near future is made clear Only simple chemical processes are discussed and very little mathematics is used Both the pleasure and use of this field of research are revealed to the interested reader The expected readership is made of high school students non chemistry major freshmen and general audience with an interest in chemistry

The real aim of this book is to prompt the reader to wonder **The Journal of Physical Chemistry** ,1903

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