

Chemistry 12 Worksheet 2-3

[KEY]

Calculations Involving the Equilibrium Constant K_{eq}

1. Given the equilibrium equation below:



If, at equilibrium, the concentrations are as follows:

$$[A_2] = 3.45 \text{ M}, \quad [B_2] = 5.67 \text{ M} \quad \text{and} \quad [AB] = 0.67 \text{ M} \quad (2 \text{ SD})$$

- a) Write the expression for the equilibrium constant,
- K_{eq}

$$K_{eq} = \frac{[AB]^2}{[A_2][B_2]}$$

- b) Find the value of the equilibrium constant,
- K_{eq}
- at the temperature that the experiment was done.

$$K_{eq} = \frac{(0.67)^2}{(3.45)(5.67)} = 0.023$$

$$\text{Answer } K_{eq} = 0.023$$

2. Given the equilibrium equation:

at a temperature of 50°C , it is found that when equilibrium is reached that:

$$[X_2] = 0.37 \text{ M}, \quad [Y_2] = 0.53 \text{ M} \quad \text{and} \quad [XY_3] = 0.090 \text{ M} \quad (\text{All } 2 \text{ SD})$$

- a) Write the equilibrium constant expression (
- K_{eq}
-)

$$K_{eq} = \frac{[XY_3]^2}{[X_2][Y_2]^3}$$

- b) Calculate the value of
- K_{eq}
- at
- 50°C
- .

$$K_{eq} = \frac{(0.090)^2}{(0.37)(0.53)^3} = 0.15$$

$$\text{Answer } K_{eq} = 0.15$$

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The Pearson Guide to Objective Chemistry for the AIEEE Singhal Atul, 2010-09

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, **The Pearson Guide to Physical Chemistry for the IIT JEE** Singhal, Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation The Old Riddle and the Newest Answer John Gerard, 1904 **Fundamentals of Physical**

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