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> Coursebook answers

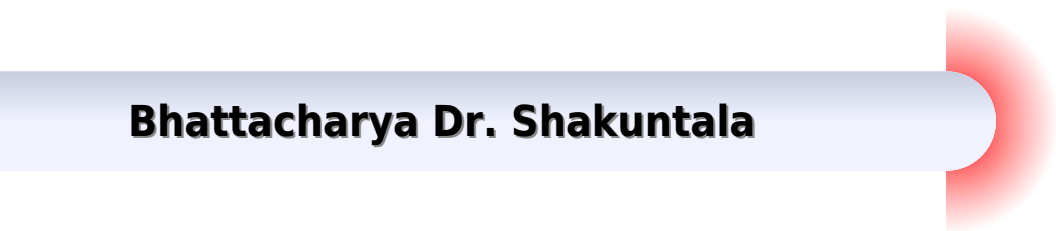
Chapter 7

Self-assessment questions

- 1 **a**
 - i H_2
 - ii CO
 - iii Mg**b**
 - i I_2O_5
 - ii SO_2
 - iii CH_3CH_3
- 2 **a**
 - i $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$ reduction
 $2\text{I}^- \rightarrow \text{I}_2 + 2\text{e}^-$ oxidation
 - ii $2\text{Mg} \rightarrow 2\text{Mg}^{2+} + 4\text{e}^-$ (or
 $\text{Mg} \rightarrow \text{Mg}^{2+} + 2\text{e}^-$) oxidation
 $\text{O}_2 + 4\text{e}^- \rightarrow 2\text{O}^{2-}$ reduction
 - iii $4\text{Fe} \rightarrow 4\text{Fe}^{2+} + 12\text{e}^-$ (or
 $\text{Fe} \rightarrow \text{Fe}^{2+} + 2\text{e}^-$) oxidation
 $3\text{O}_2 + 12\text{e}^- \rightarrow 6\text{O}^{2-}$ (or
 $\text{O}_2 + 4\text{e}^- \rightarrow 2\text{O}^{2-}$) reduction**b** $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$
- 3 **a** +5
b +6
c -2
d +3
e -3
f +3
g +4
- 4 **a**
 - i 0 to -1 = -1 (reduction)
 - ii -3 to 0 = +3 (oxidation)
 - iii +3 to +5 = +2 (oxidation)**b**
 - i Oxidising agent is Br_2 ; reducing agent is I^-
 - iii Reducing agent is As_2O_3 ; oxidising agent is I_2
- 5 **a** sodium sulfate(IV)
b sodium sulfate(VI)
c iron(II) nitrate(V) or iron(II) nitrate
d iron(III) nitrate(V) or iron(III) nitrate
e iron(II) sulfate(VI) or iron(II) sulfate
f copper(I) oxide
g sulfuric(IV) acid
h manganese(VII) oxide
- 6 **a** NaClO
b Fe_2O_3
c KNO_3
d PCl_5
- 7 **a** $\text{H}_2\text{SO}_4 + 6\text{HI} \rightarrow \text{S} + 3\text{I}_2 + 4\text{H}_2\text{O}$
b $2\text{HBr} + \text{H}_2\text{SO}_4 \rightarrow \text{Br}_2 + \text{SO}_2 + 2\text{H}_2\text{O}$
c $2\text{V}^{3+} + \text{I}_2 + 2\text{H}_2\text{O} \rightarrow 2\text{VO}^{2+} + 2\text{I}^- + 4\text{H}^+$
- 8 **B**

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Bhattacharya Dr. Shakuntala



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1. Understanding the eBook Pearson Chemistry Chapter 7 Self Check Activity Answers
 - The Rise of Digital Reading Pearson Chemistry Chapter 7 Self Check Activity Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Chapter 7 Self Check Activity Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Chapter 7 Self Check Activity Answers
 - Personalized Recommendations

- Pearson Chemistry Chapter 7 Self Check Activity Answers User Reviews and Ratings
- Pearson Chemistry Chapter 7 Self Check Activity Answers and Bestseller Lists
- 5. Accessing Pearson Chemistry Chapter 7 Self Check Activity Answers Free and Paid eBooks
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 - Adjustable Fonts and Text Sizes of Pearson Chemistry Chapter 7 Self Check Activity Answers
 - Highlighting and Note-Taking Pearson Chemistry Chapter 7 Self Check Activity Answers
 - Interactive Elements Pearson Chemistry Chapter 7 Self Check Activity Answers
- 8. Staying Engaged with Pearson Chemistry Chapter 7 Self Check Activity Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Chapter 7 Self Check Activity Answers
- 9. Balancing eBooks and Physical Books Pearson Chemistry Chapter 7 Self Check Activity Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Chapter 7 Self Check Activity Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chemistry Chapter 7 Self Check Activity Answers
 - Setting Reading Goals Pearson Chemistry Chapter 7 Self Check Activity Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chemistry Chapter 7 Self Check Activity Answers
 - Fact-Checking eBook Content of Pearson Chemistry Chapter 7 Self Check Activity Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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