

## REVIEW QUESTIONS

- Identify the reactant that acts as an acid in each of the following reactions.
  - $\text{NH}_4^+(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{NH}_3(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$
  - $\text{NH}_3(\text{g}) + \text{HCl}(\text{g}) \rightarrow \text{NH}_4\text{Cl}(\text{s})$
  - $\text{HCO}_3^-(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_3^{2-}(\text{aq})$
  - $\text{SO}_4^{2-}(\text{aq}) + \text{H}_3\text{O}^+(\text{aq}) \rightarrow \text{HSO}_4^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$
  - $\text{CO}_3^{2-}(\text{aq}) + \text{CH}_3\text{COOH}(\text{aq}) \rightarrow \text{HCO}_3^-(\text{aq}) + \text{CH}_3\text{COO}^-(\text{aq})$
- Write balanced equations to show that in water:
  - $\text{PO}_4^{3-}$  acts as a base
  - $\text{H}_2\text{PO}_4^-$  acts as an amphiprotic substance
  - $\text{H}_2\text{S}$  acts as an acid.
- Write the formula for the conjugate of:
  - the acid  $\text{HCl}$
  - the base  $\text{OH}^-$
  - the base  $\text{O}^{2-}$
  - $\text{HSO}_4^-$  when it acts as an acid.
- Using suitable examples, distinguish between:
  - a diprotic and an amphiprotic substance
  - a strong and a weak acid.
- Draw a structural formula of the monoprotic ethanoic acid molecule. Identify which proton is donated in an
- Which of the following species is both diprotic and amphiprotic?
  - $\text{HCl}$
  - $\text{HSO}_4^-$
  - $\text{H}_2\text{SO}_3$
  - $\text{H}_2\text{PO}_4^-$
- Complete and balance the following reactions:
  - $\text{Fe}(\text{s}) + 2\text{HF}(\text{aq}) \rightarrow$
  - $\text{HClO}_4(\text{aq}) + \text{LiOH}(\text{aq}) \rightarrow$
  - $\text{HNO}_3(\text{aq}) + \text{KHCO}_3(\text{aq}) \rightarrow$
  - $\text{Li}_2\text{CO}_3(\text{aq}) + 2\text{CH}_3\text{COOH}(\text{aq}) \rightarrow$
- For each of the chemical equations in Question 10, write the corresponding net ionic equation.
- What volume of  $0.0010 \text{ mol L}^{-1}$  potassium hydroxide

# Pearson Chemistry Chapters 12 Review Answers

**ML Morrison**



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