

Actual & Percent Yield



$$\frac{452 \text{g Fe}}{1} \times \frac{1 \text{mol Fe}}{55.85 \text{g}} \times \frac{1 \text{mol Fe}_3\text{O}_4}{3 \text{mol Fe}} \times \frac{231.55 \text{g}}{1 \text{mol Fe}_3\text{O}_4}$$

$$= 62.5 \text{g Fe}_3\text{O}_4$$

Pearson Chemistry Limiting Reagent And Percent Yield Workbook

Nivaldo J. Tro



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Teaching Science for Understanding James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities *Students Study Guide* Fredeen, Pearson, 2003-07 Bulletin of the Chemical Society of Japan Nihon Kagakkai, 1996 **Limiting Reactant and Percent** Lifeliqe, 2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction **Introductory Chemistry** Nivaldo J. Tro, 2006

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