

## Answers

- 55  $M = 133 \text{ g mol}^{-1}$  so gas is Xe
- 56  $90.4 \text{ g mol}^{-1}$
- 57 Helium
- 58  $311 \text{ dm}^3$
- 59 empirical formula and molecular formula =  $\text{SO}_2$
- 60 At higher altitude the external pressure is less. As the air in the tyre expands on heating (due to friction with the road surface), the internal pressure increases.
- 61 (a) Particles are in constant random motion and collide with each other and with the walls of the container in perfectly elastic collisions. The kinetic energy of the particles increases with temperature. There are no inter-particle forces and the volume of the particles is negligible relative to the volume of the gas.
- (b) At low temperature, the particles have lower kinetic energy, which favours the formation of inter-particle forces and reduces gas pressure.  $\frac{PV}{nRT} < 1$
- 62  $\text{NH}_3$  shows greater deviation than  $\text{CH}_4$  due to stronger intermolecular attractions, especially at low temperature.
- 63 B
- 64  $2.81 \text{ g}$
- 65  $4.93 \text{ g}$
- 66  $0.0100 \text{ mol}$
- 67  $0.400 \text{ mol dm}^{-3}$
- 68  $3.1 \text{ cm}^3$
- 69  $0.178 \text{ mol dm}^{-3}$
- 70  $0.0220 \text{ mol dm}^{-3}$ ,  $0.0602\% \text{ HCl}$
- 71  $0.106 \text{ mol dm}^{-3} \text{ Na}_2\text{SO}_4$  and  $0.115 \text{ mol dm}^{-3} \text{ Pb}(\text{NO}_3)_2$ ; assume no side reactions, all  $\text{PbSO}_4$  precipitates
- 72 1217 tonnes

4

- 73  $52\% \text{ NH}_3$  by mass; assuming no side reactions occur and gases behave as ideal gases
- 74  $3.20 \times 10^2 \text{ kg}$
- 75  $0.225 \text{ mol dm}^{-3}$  (or round to  $0.23 \text{ mol dm}^{-3}$ )
- 76  $[\text{Na}_2\text{CO}_3] = \frac{VP}{x} \text{ mol dm}^{-3}$   
 $[\text{NaHCO}_3] = \frac{V(P - P)}{x} \text{ mol dm}^{-3}$
- 77  $100 \text{ mol O}_2$ ;  $3.2 \text{ kg}$

## Challenge yourself

- In cold climates, temperature may approach or go below the boiling point of butane so the butane stays liquid even when it is released from the pressure it is under when stored in its canister. This makes it ineffective as a fuel.
- $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ ,  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ ,  $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$
- N = 16%, P = 22%, K = 17%
- Many reactions with 'useless' by-products could have high stoichiometric yield under optimum conditions, but low atom economy, for example, methanoic acid production:  
 $2\text{NaCOOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{HCOOH} + \text{Na}_2\text{SO}_4$   
 For 100% conversion with stoichiometric reactants, the yield = 100%.  

$$\text{atom economy} = \frac{2 \times 46.03}{(2 \times 68.01) + 98.08} \times 100\% = 39\%$$
- $2\text{NaN}_3(\text{s}) \rightarrow 2\text{Na}(\text{s}) + 3\text{N}_2(\text{g})$   
 $10\text{Na}(\text{s}) + 2\text{KNO}_3(\text{s}) \rightarrow \text{K}_2\text{O}(\text{s}) + 5\text{Na}_2\text{O}(\text{s}) + \text{N}_2(\text{g})$   
 $\text{K}_2\text{O}(\text{s}) + \text{Na}_2\text{O}(\text{s}) + \text{SiO}_2(\text{s}) \rightarrow \text{Na}_2\text{K}_2\text{SiO}_6$  (alkaline silicate glass)
- As  $\text{NaOH}$  dissolves, the separated  $\text{Na}^+$  and  $\text{OH}^-$  ions become hydrated, i.e. they are surrounded by  $\text{H}_2\text{O}$  molecules. This involves breaking the hydrogen bonds between the  $\text{H}_2\text{O}$  molecules in pure water and allows closer packing, which reduces the volume.

# Pearson Chemistry Text With Answers

**Manisha Agrawal**



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The Old Riddle and the Newest Answer John Gerard, 1904      **Chemistry: The Central Science** Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation      **The New iPad Fully Loaded** Alan Hess, 2012-05-14 Maximize all that the new iPad has to offer with The New iPad Fully Loaded The new iPad Fully Loaded is a one of a kind resource for squeezing every bit of functionality from your new iPad Whether you re downloading content of any sort looking to get the most out of iOS5 or using your new iPad to control anything from your home stereo system to your vehicle nothing is out of the realm of possibilities As you progress through this captivating full color book veteran author Alan Hess takes you beyond the basics and show you how to fully utilize your new iPad and it s impressive features Peppered with sophisticated tips and tricks The new iPad Fully Loaded encourages you to modify your new iPad to your own specifications and teaches you never before revealed tricks so that you can truly get the most out of this amazing device Goes beyond the basics to deliver rare tips and tricks on maximizing all that the new iPad has to offer Includes tips tricks and techniques that are applicable to every version of the iPad Covers new benefits of iOS 5 Boasts updates on the latest iPad technology developments as well as coverage of both common and uncommon tasks In full color and sporting a convenient trim size The new iPad Fully Loaded is the perfect accessory to your new iPad      **The Burning Answer** Keith Barnham, 2015-05-15 Our society faces a choice We could be enjoying a sustainable lifestyle but we have chosen not to In three generations we have consumed half the oil produced by photosynthesis over eight million generations In two generations we have used half our uranium resources With threats from global warming oil depletion and nuclear disaster we are running out of options Solar power as Keith Barnham says is our necessary solution Barnham explains that the roots of solar energy lie in a little known equation  $E = hf$  an equation which was coincidentally celebrated and explained to the world by Einstein in the same year he discovered  $E = mc^2$  He alleges that the former equation has been overlooked in favor of the latter much to our detriment and Barnham is here to offer us a solution We can still turn things around and solar energy is the key In this provocative inspiring passionately argued book Keith Barnham outlines actions that any one and all of us can take to make an impact now and on future generations The Burning

Answer is a solar manifesto for the new climate aware generation and a must read for climate change skeptics **A**

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Agrawal,2021-11-01 Chemistry I is a compulsory paper for the first year Undergraduate course in Engineering Technology Syllabus of this book is strictly aligned as per model curriculum of AICTE and academic content is amalgamated with the concept of outcome based education Book covers seven topics Atomic and molecular structure Spectroscopic Technique and applications Inter molecular Forces and Potential Energy Surfaces Use of Free Energy in Chemical Equilibrium Periodic Properties Stereo chemistry Organic Reactions and Synthesis of Drug Molecules Each topic is written in easy and lucid manner Every chapter contains a set of exercise at the end of each unit to test student's comprehension Salient Features Content of the book aligned with the mapping of Course Outcomes Programs Outcomes and Unit Outcomes Book Provides lots of recent information interesting facts QR Code for E resources QR Code for use of ICT Projects group discussion etc Students and teacher centric subject materials included in book with balanced and chronological manner Figures tables chemical equations and comparative charts are inserted to improve clarity of the topics Short questions objective questions and long answer exercises are given for practice of students after every chapter Solved and unsolved problems including numerical examples are solved with systematic steps *Prentice Hall Chemistry* Antony C. Wilbraham,2008 The Big

Book of Chemistry Teacher Stories Jeff Lark, Stories from years of teaching high school chemistry **Examination**

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