Stoichiometric Calculations Problems complete on a sheet and submit to the dropbox.
 Show all working.

- 4NH3(g) + 6NO (g) → 5N2(g) + 6H2O (g): (a) How many moles of water is produce in this reaction? 4.6 B. How many grams of grams of NO was needed in the reaction? 30.01g
- Using the following equation to answer (a1 -b2): Therefore, 72.08 grams of H2O will be produced.

2 NaOH + H2SO4 ---> 2 H2O + Na2SO4

a1. How many moles of Na2S04 was produced when 1mole of H2S04 reacted with 2mOles of NaOH? when 1 mole of H2S04 reacts with 2 moles of NaOH, 1 mole of Na2S04 is produced.

b2. How many grams of H20 will be produced if 160grams of NaOH reacts with H2SO4. 72.08 grams of H2O will be produced when 160 grams of NaOH reacts with H2SO4.

# **Pearson Chemistry Work Stoichiometry**

Jean-Claude G. Bunzli, Vitalij K. Pecharsky

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