13   These calculations have used $L = 6.02 \times 10^{23}$

    (a) 7.22×10^{22} (b) 3.01×10^{24}

    (c) 1.20×10^{23}

14   0.53 mol H

15   0.250 mol

16   (a) 262.87 g mol^{-1} (b) 176.14 g mol^{-1}

    (c) 164.10 g mol^{-1} (d) 248.22 g mol^{-1}

17   189.1 g

18   1.5 mol

19   0.0074 mol Cl$^-$

20   1.83×10^{23} C atoms

21   171 g (integer value because no calculator)

22   10.0 g H_2O

23   2.0 mol N_2 > 3.0 mol NH_3 > 25.0 mol H_2 >
    1.0 mol N_2H_4

24   (a) CH   (b) CH_2O

    (c) $C_{12}H_{22}O_{11}$ (d) C_2H_6

    (e) C_2H_5 (f) CH_2O

25   $Na_2S_2O_3$

26   $CoSO_4 \cdot 7H_2O$

27   $C_3H_{10}N$

28   NH_3

29   6.94 Li

30   CdS

31   empirical formula CH; molecular formula C_6H_6

32   empirical formula H_3PO_2; molecular formula $H_6P_2O_4$

33   $C_8H_{10}N_4O_2$ for both empirical and molecular
    formulas

34   C_2H_4O

35   Let y = mass of chalk in grams.

    moles of chalk used $= \dfrac{\text{mass used}}{M_r(CaCO_3)} =$

    $\dfrac{y \text{ g}}{100.09 \text{ g mol}^{-1}}$

This is the same as the number of moles of carbon atoms used.

Therefore the number of carbon atoms used = moles of chalk $\times$ $(6.02 \times 10^{23}$ mol$-1) =$

$\dfrac{6.02 \times 10^{23} \, y}{100.09}$

36   (a) 2.50 mol   (b) 5.63 mol

    (c) 665.5 g

37   (a) $2C_4H_{10} + 13O_2 \rightarrow 8CO_2 + 10H_2O$

    (b) 1.59 g

38   4.355 kg

39   (a) $CaCO_3 \rightarrow CaO + CO_2$

    (b) 92.8%

    (c) $CaCO_3$ is the only source of CO_2; all the
       $CaCO_3$ undergoes complete decomposition;
       all CO_2 released is captured; heating does
       not cause any change in the mass of the
       other minerals present.

40   (a) 85.2 g   (b) 1.3 g H_2

41   5.23 g $C_2H_5Cl_2$

42   253.2 g theoretical $CaSO_4$; 77.9%

43   3.16 g ester

44   107 g of C_2H_6 needed

45   (a) 2.40 mol   (b) 0.0110 mol

    (c) 44 mol

46   (a) 35.65 dm^3 (b) 5.7 dm^3

47   0.652 dm^3

48   0.138 mol Br_2 and 0.156 mol Cl_2, so more
    molecules of Cl_2

49   0.113 dm^3

50   0.28 dm^3

51   90 kPa

52   16 °C

53   3.0 dm^3

54   2.8 dm^3

3

Pearson Chemistry Workbook Answers Chapter 1stoichiometry

Elissa Huddart

Pearson Chemistry Workbook Answers Chapter 1stoichiometry:

Chemistry and Industry ,1954 *The Software Encyclopedia* ,1988 *Students Study Guide* Fredeen,Pearson,2003-07 **Student Workbook and Selected Solutions Manual for Introductory Chemistry** Steve Russo,Michael Silver,2014-02-07 Features examples from each chapter learning objectives review of key concepts from the text and additional problems for student practice Also the workbook provides comprehensive answers and explanations to selected end of chapter problems from the text Provides over 200 worked examples and more than 550 practice problems and quiz questions to help students develop and practice their problem solving skills **Stoichiometry** Sydney Young,1908 *Formulation and Stoichiometry* Emil J. Margolis,2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in con trasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction Stoichiometry and Structure - Freshman Chemistry Problems and how to Solve Them - Part 1 ,1964 **Chemistry** Neil D. Jespersen,Alison Hyslop,2014-01-13 The Student Solutions Manual to accompany Chemistry The Molecular Nature of Matter 7th Edition Jespersen s Chemistry The Molecular Nature of Matter 7th Edition provides readers with the necessary practice support instruction and assessment that is required for learning and teaching the content of a General Chemistry course This text provides the forum for problem solving and concept mastery of chemical phenomena that leads to proficiency and success The Seventh Edition includes revisions to key content coverage areas and concepts and the addition of more Analyzing Solving Multi Concept problems and examples throughout the text An increased emphasis has also been placed on the intimate relationship that exists between structure at the submicroscopic molecular level and the observable macroscopic properties of matter Jespersen provides readers with a clear concise and easy to understand General Chemistry resource *Mole Concepts and Stoichiometry* Dharshi Bopegedera,2017 **Calculating in Chemistry** Cosmas C Mbuzi,2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may

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