Answers

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4

73 52% NH, by mass; assuming no side reactions occur and gases behave as ideal gases

76
$$\text{[Nia_jOO_j]} = \frac{\text{YP}}{\text{X}} \text{ mol dm}^{-1}$$

 $\text{[NiaHOO_j]} = \frac{\text{Y}(Q - P)}{\text{X}} \text{ mol dm}^{-1}$

Challenge yourself

1 In cold climates, temperature may approach or go below the boiling point of butane so the butane stays liquid even when it is released from the pressure it is under when stored in its canister. This makes it ineffective as a fuel.

4 Many reactions with "useless" by-products could have high stoichiometric yield under optimum conditions, but low atom economy, for example, methanoic acid production:

For 100% conversion with stoichiometric reactants, the yield = 100%.

atom economy =
$$\frac{2 \times 46.03}{(2 \times 68.01) + 98.08} \times 100\% = 39\%$$

- 5 2NaN,(s) → 2Na(s) + 3N,(g) 10Na(s) + 2KNO,(s) → K,O(s) + 5Na,O(s) + N,(g) K,O(s) + Na,O(s) + SiO,(s) → Na,K,SiO, (alkaline silicate glass)
- 6 As NaOH dissolves, the separated Nar and OHions become hydrated, i.e. they are surrounded by H₂O molecules. This involves breaking the hydrogen bonds between the H₂O molecules in pure water and allows closer packing, which reduces the volume.

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