

Periodic Table Puzzle

The object of this puzzle is to use the periodic properties of the elements to fill in the following chart. The code letters A to Z have been assigned to the first 26 representative elements in the periodic table. The code letters do not represent certain chemical symbols, that is, C may not be carbon, nor have the letters been assigned in alphabetical order.

Your task is to arrange these letters in the proper location in the periodic chart. To do this, use the clues given pertaining to certain members of each group.

Clues:

The following elements belong together in groups: WXYZ, STUV, ABC, DEF, GHI, JKL, MNO, and PQR.

1. Element M is not the smallest in its group.
2. Elements L & E are in the same period.
3. The atomic radius P is smaller than element R.
4. Element T has the lowest ionization energy in its group.
5. Element W is more reactive than element A and element A is more reactive than element D. All three elements are in the same period.
6. The atomic radius of H is smaller than element I.
7. Element B is the least reactive alkaline earth metal.
8. Element K has 5 valence electrons.
9. L has a lower ionization energy than K.
10. Element Q is the most reactive halogen.
11. The atomic number of element V is one more than element P.
12. Element J is the most reactive element in its group.
13. Elements C and N are from the same period and react to form CN.
14. Element O forms a -2 ion.
15. Element Z is the most reactive metal in its period.
16. The atomic number of Z is 16 more than the atomic number of Y.
17. Element E reacts with element P to form the ionic compound EP_3 .
18. Element U has the smallest radius in its period, but not in its group.
19. Element G has the highest ionization energy in its group.

Pearson Education Logic Problem Periodic Table Answer

Rachel S Tattersall



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