

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

1. Calculate the number of moles of potassium chlorate, KClO_3 (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

Equation:	$2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$			1.2 moles KClO_3 needed
Before:	1.2	0	0	
Change:	-1.2	+1.2	+1.8	
After:	0	1.2	1.8	

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equation:	$\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$				0.200 moles HCl needed
Before:	0.100	?	0	0	
Change:	-0.100	-0.200 +	0.100 +	0.100	
After:	0	?	0.100	0.100	

3. Ethane, C_2H_6 reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	$2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$			
Before:	2.20	xs	0	0
Change:	-2.20	- 7.70	+ 4.40	+ 6.60
After:	0	xs	4.40	6.60

O_2 needed: $7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$

Equation:	$\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$			
Before:	0.0040	XS	0	0
Change:	- 0.0040	- 0.0080	+ 0.0040	+ 0.0080
After:	0	XS	0.0040	0.0080

NaNO_3 produced: $0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$

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**Richard Longden Boole, Alyn G.
McFarland**



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1. Understanding the eBook Pearson Education Science Workbook Answers Chemical
 - The Rise of Digital Reading Pearson Education Science Workbook Answers Chemical
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Education Science Workbook Answers Chemical
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Education Science Workbook Answers Chemical
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Education Science Workbook Answers Chemical
 - Personalized Recommendations
 - Pearson Education Science Workbook Answers Chemical User Reviews and Ratings

- Pearson Education Science Workbook Answers Chemical and Bestseller Lists
- 5. Accessing Pearson Education Science Workbook Answers Chemical Free and Paid eBooks
 - Pearson Education Science Workbook Answers Chemical Public Domain eBooks
 - Pearson Education Science Workbook Answers Chemical eBook Subscription Services
 - Pearson Education Science Workbook Answers Chemical Budget-Friendly Options
- 6. Navigating Pearson Education Science Workbook Answers Chemical eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Education Science Workbook Answers Chemical Compatibility with Devices
 - Pearson Education Science Workbook Answers Chemical Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Education Science Workbook Answers Chemical
 - Highlighting and Note-Taking Pearson Education Science Workbook Answers Chemical
 - Interactive Elements Pearson Education Science Workbook Answers Chemical
- 8. Staying Engaged with Pearson Education Science Workbook Answers Chemical
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Education Science Workbook Answers Chemical
- 9. Balancing eBooks and Physical Books Pearson Education Science Workbook Answers Chemical
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Education Science Workbook Answers Chemical
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Education Science Workbook Answers Chemical
 - Setting Reading Goals Pearson Education Science Workbook Answers Chemical
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Education Science Workbook Answers Chemical
 - Fact-Checking eBook Content of Pearson Education Science Workbook Answers Chemical
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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