

# Stoichiometry: Practice Assessment

Name: \_\_\_\_\_

**5888** - Using a balanced equation, given the grams of reactant, calc. the grams of product



Balance the equation provided. Calculate the number of moles of sodium bromide that would be produced if 25.5 moles of gallium (III) bromide are used (with an excess of sodium sulfite)

$$25.5\text{mol GaBr}_3 \times 6\text{mol NaBr}/2\text{mol GaBr}_3 = 75.75\text{mol NaBr}$$

**2)** Tin (II) oxide reacts with nitrogen trifluoride to produce tin (II) fluoride and dinitrogen trioxide. Using the following reaction, calculate the grams produced of each product from 50.0 grams of nitrogen trifluoride.



$$50\text{g NF}_3 \times 1\text{mol NF}_3 / 71.001\text{g NF}_3 = 0.704\text{mol NF}_3 \times 2\text{mol SnF}_2 / 2\text{mol NF}_3 \times 156.69\text{g SnF}_2 / 1\text{mol SnF}_2 = 165.5\text{g SnF}_2$$

**SnF<sub>2</sub>**

$$= 0.704\text{mol NF}_3 \times 1\text{mol N}_2\text{O}_3 / 2\text{mol NF}_3 \times 76.014\text{g N}_2\text{O}_3 / 1\text{mol N}_2\text{O}_3 =$$

**26.75g N<sub>2</sub>O<sub>3</sub>**

**3)** A piece of copper wire is placed in silver nitrate. Silver and aqueous copper (II) nitrate is produced. The initial mass of the copper before the reaction was 2.74g and the mass remaining after the reaction was 2.32g. What is the mass of silver produced?



$$2.74\text{g} - 2.32\text{g} = 0.42\text{g}$$

$$0.42\text{g Cu} \times 1\text{mol Cu} / 63.55\text{g Cu} \times 2\text{mol Ag} / 1\text{mol Cu} \times 107.87\text{g Ag} / 1\text{mol Ag} = 1.43\text{g Ag}$$

**5889** - After calculating theoretical yield in grams, calculate the percent yield. Given experimental data, determine the limiting reactant and the excess reactant

**a)** Through experimentation, you observe the reaction of iron (III) phosphate with sodium sulfate to make iron (III) sulfate and sodium phosphate. You began the experiment with 25g of iron (III) phosphate and 182g of sodium sulfate. How many grams of iron (III) sulfate should you have been able to produce? What is the percent yield if 18.5g of iron (III) sulfate are actually made?



$$25\text{g FePO}_4 \times 1\text{mol FePO}_4 / 159.815\text{g FePO}_4 = 0.156\text{mol FePO}_4 / 2\text{mol per rxn} = 0.083\text{rxns (LR)}$$

$$182\text{g Na}_2\text{SO}_4 \times 1\text{mol Na}_2\text{SO}_4 / 142.04\text{g Na}_2\text{SO}_4 = 1.28\text{mol Na}_2\text{SO}_4 / 3\text{mol per rxn} = 0.427\text{rxns}$$

$$0.083\text{mol FePO}_4 \times 1\text{mol Fe}_2(\text{SO}_4)_3 / 2\text{mol FePO}_4 \times 399.873\text{g Fe}_2(\text{SO}_4)_3 / 1\text{mol Fe}_2(\text{SO}_4)_3 = 33.19\text{g}$$

$$(18.5\text{g} / 33.19\text{g}) \times 100 = 56\%$$

# Pearson Education Stoichiometry Answer Key

**Raffaella Di Napoli**



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