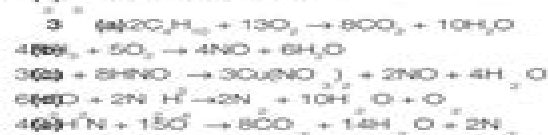
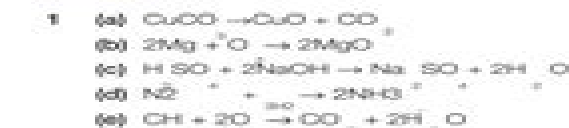


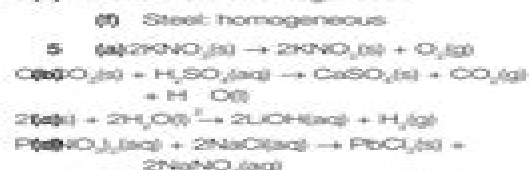
Answers

Chapter 1

Exercises



4 (a) Sand and water: heterogeneous
 (b) Sugar and water: homogeneous
 (c) Oil and iron filings: heterogeneous
 (d) Endol and water: homogeneous



In the answers, exercises and text there is inconsistency about whether the water formed in combustion reactions is gas or liquid. At room temperature it will ultimately finish up as a liquid, as shown here.

6 X has diffused more quickly, so it must be a lighter gas. Its particles have greater velocity than the particles of Y at the same temperature. (Note though that they will both have the same value for average kinetic energy.)

7 From the kinetic molecular theory we would expect a solid to be more dense than its liquid, and therefore that ice would sink in water.

8 Bubbles will be present through the volume of the liquid. A brown gas is visible above the brown liquid. As the two states are at the same temperature, the particles have the same average kinetic energy and are moving at the same speed. The inter-particle distances in the gas are significantly larger than those in the liquid.

9 At certain conditions of low temperature and low humidity, snow changes directly to water vapour by sublimation, without going through the liquid phase.

10 Steam will condense on the skin, releasing energy as it forms liquid at the same temperature (E-D on Figure 1.4). This is additional to the energy released when both the boiling water and the condensed steam cool on the surface of the skin.

11 B

12 AW 01_EX_12

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Rex M. Heyworth



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Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Darius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

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Analysis of Kinetic Data E. S. Swinbourne, 2013-04-09 Data analysis is important from two points of view first it enables a large mass of information to be reduced to a reasonable compass and second it assists in the interpretation of experimental results against some framework of theory The purpose of this text is to provide a practical introduction to numerical methods of data analysis which have application in the field of experimental chemical kinetics Recognizing that kinetic data have many features in common with data derived from other sources I have considered it appropriate to discuss a selection of general methods of data analysis in the early chapters of the text It is the author's experience that an outline of these methods is not always easy to locate in summary form and that their usefulness is often not sufficiently appreciated Inclusion of these methods in the early chapters has been aimed at simplifying discussion in the later chapters which are more particularly concerned with kinetic systems By the provision of a number of worked examples and problems it is hoped that the reader will develop a feeling for the range of methods available and for their relative merits Throughout the text the mathematical treatment has been kept relatively simple lengthy proofs being avoided I have preferred to indicate the sense and usefulness of the various methods rather than to justify them on strict mathematical grounds

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