

Chemistry Chapter 10 Test

True/False

Indicate whether the sentence or statement is true or false.

1. KMT stands for Kinetic Mole Theory.
2. One of the assumptions in the KMT is that the particles are spread very far apart.
3. Collisions between gas particles and between particles and container walls are elastic collisions.
4. The collisions that gas particles undergo will have a loss of energy according to the KMT.
5. According to the KMT there is an attraction between particles.
6. Gas particles are in constant, random, and rapid motion.
7. Gas particles exhibit the property of expansion.
8. Gas particles have a high density and therefore a low compressibility.
9. The spontaneous mixing of gas particles is called effusion.
10. Ideal gases do not really exist.
11. Particles colliding with walls of a container create pressure.
12. There are 760 torr in 1 atm.
13. There are 101.3 mm Hg in 1 atm.
14. Boyle's law deals with volume and temperature.
15. The combined gas law is composed of volume, pressure, and temperature.

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

16. The standard molar volume of a gas at STP is
a. 22.4 L
b. 22.4 L/mol
c. g-mol wt/22.4 L
d. 22.4 L/g-mol wt

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L Reisser



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