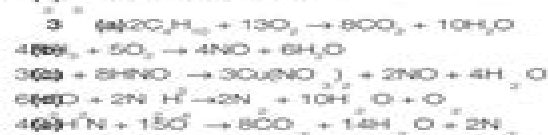
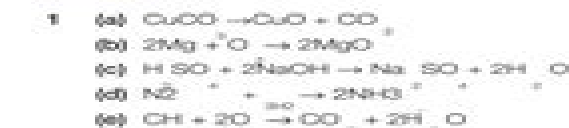


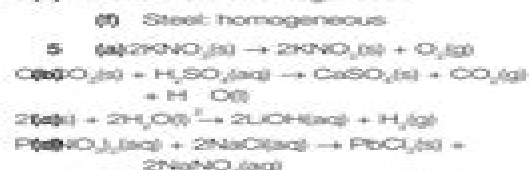
Answers

Chapter 1

Exercises



4 (a) Sand and water: heterogeneous
 (b) Sugar and water: homogeneous
 (c) Oil and iron filings: heterogeneous
 (d) Endol and water: homogeneous



In the answers, exercises and text there is inconsistency about whether the water formed in combustion reactions is gas or liquid. At room temperature it will ultimately finish up as a liquid, as shown here.

6 X has diffused more quickly, so it must be a lighter gas. Its particles have greater velocity than the particles of Y at the same temperature. (Note though that they will both have the same value for average kinetic energy.)

7 From the kinetic molecular theory we would expect a solid to be more dense than its liquid, and therefore that ice would sink in water.

8 Bubbles will be present through the volume of the liquid. A brown gas is visible above the brown liquid. As the two states are at the same temperature, the particles have the same average kinetic energy and are moving at the same speed. The inter-particle distances in the gas are significantly larger than those in the liquid.

9 At certain conditions of low temperature and low humidity, snow changes directly to water vapour by sublimation, without going through the liquid phase.

10 Steam will condense on the skin, releasing energy as it forms liquid at the same temperature (E–D on Figure 1.4). This is additional to the energy released when both the boiling water and the condensed steam cool on the surface of the skin.

11 B

12 AW 01_EX_12

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N Noddings



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1. Understanding the eBook Pearson Chemistry Lesson Check Answers 14
 - The Rise of Digital Reading Pearson Chemistry Lesson Check Answers 14
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Lesson Check Answers 14
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Lesson Check Answers 14
 - Personalized Recommendations
 - Pearson Chemistry Lesson Check Answers 14 User Reviews and Ratings
 - Pearson Chemistry Lesson Check Answers 14 and Bestseller Lists
5. Accessing Pearson Chemistry Lesson Check Answers 14 Free and Paid eBooks
 - Pearson Chemistry Lesson Check Answers 14 Public Domain eBooks
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 - Pearson Chemistry Lesson Check Answers 14 Enhanced eBook Features
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 - Highlighting and Note-Taking Pearson Chemistry Lesson Check Answers 14
 - Interactive Elements Pearson Chemistry Lesson Check Answers 14
8. Staying Engaged with Pearson Chemistry Lesson Check Answers 14

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- Participating in Virtual Book Clubs
- Following Authors and Publishers Pearson Chemistry Lesson Check Answers 14
- 9. Balancing eBooks and Physical Books Pearson Chemistry Lesson Check Answers 14
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Lesson Check Answers 14
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chemistry Lesson Check Answers 14
 - Setting Reading Goals Pearson Chemistry Lesson Check Answers 14
 - Carving Out Dedicated Reading Time
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 - Fact-Checking eBook Content of Pearson Chemistry Lesson Check Answers 14
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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