

## Chemistry 30

### Thermochemistry Review

1. Which of the following statements is **true** for an endothermic reaction?
- A.  $\Delta H_r$  for the reaction is negative.
  - B. Energy is a product in the chemical equation.
  - C. The temperature of the surroundings increases.
  - D. The potential energy of the products is greater than the potential energy of the reactants.

*Use the following information to answer the next question.*



2. The  $\Delta H_r$  value for the reverse reaction would be
- A. +100 kJ
  - B. +50.0 kJ
  - C. -50.0 kJ
  - D. -100 kJ

#### Numerical Response

1. Benzene is a gasoline additive. The heat of formation for 1.00 mol of  $\text{C}_6\text{H}_6\text{(l)}$  is \_\_\_\_\_ kJ.

(Record your **three-digit answer** in the numerical-response section on the answer sheet)

#### Numerical Response

2. An outdoor mercury thermometer contains 1.02 g of mercury. On a winter day, the thermometer records a change in temperature from  $-38.0^\circ\text{C}$  to  $-13.0^\circ\text{C}$ . The mercury absorbed \_\_\_\_\_ J of energy. The heat capacity of mercury is  $0.140 \text{ J/g}^\circ\text{C}$ .

(Record your **three-digit answer** in the numerical-response section on the answer sheet)

# Pearson Chemistry Section Review Answers

## Thermochemistry

**RM Cervero**



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