

13 These calculations have used $L = 6.02 \times 10^{23}$

(a) 7.22×10^{22} (b) 3.01×10^{23}

(c) 1.20×10^{23}

14 0.53 mol H

15 0.250 mol

16 (a) 262.87 g mol⁻¹ (b) 176.14 g mol⁻¹

(c) 164.10 g mol⁻¹ (d) 248.22 g mol⁻¹

17 189.1 g

18 1.5 mol

19 0.0074 mol Cl⁻

20 1.83×10^{23} C atoms

21 171 g (integer value because no calculator)

22 10.0 g H₂O

23 2.0 mol N₂ > 3.0 mol NH₃ > 25.0 mol H₂ > 1.0 mol N₂H₄

24 (a) CH (b) CH₂O

(c) C₁₀H₁₂O₁₁ (d) C₃H₆

(e) C₃H₆ (f) CH₂O

25 Na₂S₂O₅

26 CoSO₄·7H₂O

27 C₁₁H₁₅N

28 NH₃

29 6.94 L

30 CdS

31 empirical formula CH; molecular formula C₇H₈

32 empirical formula H₂PO₃; molecular formula H₄P₂O₆

33 C₁₀H₁₂N₂P₂O₁₁ for both empirical and molecular formulas

34 C₂H₃O

35 Let y = mass of chalk in grams.

moles of chalk used = $\frac{\text{mass used}}{M_r(\text{CaCO}_3)} \times$

$y \frac{\text{g}}{100.09 \text{ g mol}^{-1}}$

This is the same as the number of moles of carbon atoms used.

Therefore the number of carbon atoms used =

moles of chalk \times (6.02 \times 10²³ mol⁻¹) =

$\frac{6.02 \times 10^{23} y}{100.09}$

36 (a) 2.50 mol

(b) 5.63 mol

(c) 665.5 g

37 (a) $2\text{C}_2\text{H}_6 + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$

(b) 1.59 g

38 4.355 kg

39 (a) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$

(b) 92.8%

(c) CaCO₃ is the only source of CO₂; all the CaCO₃ undergoes complete decomposition; all CO₂ released is captured; heating does not cause any change in the mass of the other minerals present.

40 (a) 65.2 g

(b) 1.3 g H₂

41 5.23 g C₂H₅Cl₂

42 253.2 g theoretical CaSO₄; 77.9%

43 3.16 g ester

44 107 g of C₆H₆ needed

45 (a) 2.40 mol

(b) 0.0110 mol

(c) 44 mol

46 (a) 35.65 dm³

(b) 5.7 dm³

47 0.652 dm³

48 0.138 mol Br₂ and 0.156 mol Cl₂, so more molecules of Cl₂

49 0.113 dm³

50 0.28 dm³

51 90 kPa

52 16 °C

53 3.0 dm³

54 2.8 dm³

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SA Adler



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The Publishers' Circular and Booksellers' Record of British and Foreign Literature ,1901 *Technical Book Review Index* ,1961 *Chemistry-I (with Lab Manual)* Manisha Agarwal,2021-01-01 Chemistry I is a compulsory paper for the first year Undergraduate course in Engineering Technology Syllabus of this book is strictly aligned as per model curriculum of AICTE and academic content is amalgamated with the concept of outcome based education **An Introduction to Chemical Metallurgy** Roger Hill Parker,1967 **Clinical Chemistry** Michael L. Bishop,Janet L. Duben-Engelkirk,Edward P. Fody,1992 **Chemistry and Industry** ,1976

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