

Stoichiometry Presentation

mole ratios, mole $\Rightarrow$ mole, mole $\Rightarrow$ mass, mass $\Rightarrow$ mass, limiting reactants, percent yield, theoretical yield

Mass-to-Mass Conversions

- We can only use 3 things:
 1. mole ratio
 2. molar mass
 3. known



- STEPS:**
 1. write the balanced chemical equation
 2. use a **molar mass** to convert from mass to moles of known
 3. use a **mole ratio** to convert from moles of known to moles of unknown
 4. use a **molar mass** to convert from moles of unknown to mass of unknown

Mole-to-Mole Conversions

- You can only use 2 things:
 1. mole ratios
 2. known

Know	Want
	Don't

- STEPS:**
 1. write the balanced chemical equation
 2. use a **mole ratio** to convert from moles to moles of unknown



$$\begin{array}{c|c|c} 1.25 \text{ mol-Cl}_2 & 2 \text{ mol-NaCl} & 58.44 \text{ g NaCl} \\ \hline & 1 \text{ mol-Cl}_2 & 1 \text{ mol-NaCl} \end{array}$$

$$= (1.25) \times (2) \times (58.44 \text{ g NaCl})$$

$$= 146.1 \text{ g NaCl}$$

Chemistry

Limiting Reactant

- Chocolate Chip Cookie Recipe
 - 1 cup butter, softened
 - 1 cup white sugar
 - 1 cup packed brown sugar
 - 2 large eggs
 - 2 teaspoons vanilla extract
 - 1 teaspoon baking soda
 - 2 teaspoons hot water
 - ½ teaspoon salt
 - 3 cups all-purpose flour
 - 2 cups semisweet chocolate chips



4 cups flour



4 Eggs

Out of eggs and flour, which is limiting?

Percent Yield

- percent yield:** measure of the efficiency of a chemical reaction
 - no reaction is 100% efficient
- theoretical yield:** maximum amount of product that could be produced from a given amount of reactant
- actual yield:** amount of product produced when the chemical reaction is carried out in an experiment



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