

# Chemistry

## Unit 2

### Practice Exam 1 Answers

#### Section A—Multiple-choice questions

##### Question 1

- B** The high electronegativity difference between hydrogen and oxygen means that the O—H bond in water is highly polar, enabling hydrogen bonds to be formed between nearby water molecules.

Sulfur is less electronegative than oxygen. The polarity of the H—S bond in  $\text{H}_2\text{S}$  is much less than the polarity of the O—H bond in water. The attraction between  $\text{H}_2\text{S}$  molecules is described as a dipole-dipole interaction.

The bond between hydrogen and Te is non-polar. The only interactions between molecules of  $\text{H}_2\text{Te}$  are dispersion forces brought about through the formation of instantaneous dipoles.

##### Question 2

- D** The formation of a solution involves breaking bonds between solute particles, breaking intermolecular bonds between solvent molecules, and the formation of bonds between solvent molecules and solute particles.

##### Question 3

- C**  $n(\text{Na}_2\text{CO}_3) = cV = 0.150 \times \frac{300}{1000} = 0.0450 \text{ mol}$   
 $m(\text{Na}_2\text{CO}_3) = mM = 0.0450 \times 106.0 = 4.77 \text{ g}$

##### Question 4

- C** Water is a poor conductor of electricity due to the small number of hydronium ions and hydroxide ions that are produced by the self-ionisation reaction of water.

##### Question 5

- A** Nitric acid,  $\text{HNO}_3$ , is a strong acid because it is completely ionised in water. Ethanoic acid,  $\text{CH}_3\text{COOH}$ , is a weak acid as it is partly ionised in water. A 0.1 M solution is considered to be dilute and 6 M concentrated.

##### Question 6

- B** When the acid  $\text{H}_2\text{PO}_4^-$  donates a proton, it forms the conjugate base  $\text{HPO}_4^{2-}$ .

##### Question 7

- A** pH is defined as  $\text{pH} = -\log_{10}[\text{H}_3\text{O}^+]$ . Since HCl is a strong acid and completely ionised in solution, it can be assumed that in each beaker  $[\text{HCl}] = [\text{H}_3\text{O}^+]$ . pH is only dependent on the  $[\text{H}_3\text{O}^+]$ , not the volume of solution. Because the definition involves the negative logarithm, the pH increases as the  $[\text{H}_3\text{O}^+]$  decreases. The solution that has the lowest  $[\text{H}_3\text{O}^+]$  will have the highest pH.

# Pearson Education Chemistry Packet Answers

**Karen C. Timberlake**



## **Pearson Education Chemistry Packet Answers:**

**Young Scientist Series ICSE Chemistry Work Book 8 ,      Young Scientist Series ICSE Chemistry Work Book 6**  
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      **General, Organic, and Biological Chemistry Study Guide and Selected Solutions** Karen C. Timberlake,2001-11  
Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with

answers and mastery exams The guide also contains complete solutions to odd numbered problems

**Organic Chemistry Workbook** Pierre Vogel, Kendall N. Houk, 2019-08-20 Provides references and answers to every question presented in the primary Organic Chemistry textbook Successfully achieving chemical reactions in organic chemistry requires a solid background in physical chemistry Knowledge of chemical equilibria thermodynamics reaction rates reaction mechanisms and molecular orbital theory is essential for students chemists and chemical engineers The Organic Chemistry presents the tools and models required to understand organic synthesis and enables the efficient planning of chemical reactions This volume Organic Chemistry Theory Reactivity and Mechanisms in Modern Synthesis Workbook complements the primary textbook supplying the complete calculated solutions to more than 800 questions on topics such as thermochemistry pericyclic reactions organic photochemistry catalytic reactions and more This companion workbook is indispensable for those seeking clear in depth instruction on this challenging subject Written by prominent experts in the field of organic chemistry this book Works side by side with the primary Organic Chemistry textbook Includes chapter introductions and re stated questions to enhance efficiency Features clear illustrations tables and figures Strengthens reader s comprehension of key areas of knowledge Organic Chemistry Theory Reactivity and Mechanisms in Modern Synthesis Workbook is a must have resource for anyone using the primary textbook

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