

8

COVALENT BONDING

Chapter Test B

A. Matching

Match each term in Column B with the correct description in Column A. Write the letter of the correct term on the line.

Column A	Column B
_____ 1. the mixing of several atomic orbitals to form the same number of equivalent hybrid orbitals	a. molecular compound
_____ 2. two or more vald electron dot formulas that can be written for the same molecule	b. VSEPR theory
_____ 3. a chemical formula that shows the arrangement of atoms in molecules and polyatomic ions	c. van der Waals forces
_____ 4. the weakest attractions that exist between molecules	d. single covalent bond
_____ 5. because electron pairs repel, molecules adjust their shapes so that the valence electron pairs are as far apart as possible	e. resonance structures
_____ 6. a covalent bond in which one pair of electrons is shared between two atoms	f. polar molecule
_____ 7. a molecule in which one end is slightly negative and the other end is slightly positive	g. hybridization
_____ 8. a covalent bond formed between two different atoms in which the bonding electrons are shared unequally	h. structural formula
_____ 9. an electrically neutral group of atoms joined together by covalent bonds	i. polar covalent bond
_____ 10. a covalent bond in which one atom contributes both bonding electrons	j. coordinate covalent bond

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