Chemistry Student Edition - Basic Answer Key

Chapter 17: Thermochemistry

Heat Flow

Questions

- 1. What is true of all chemical processes?
- How can you distinguish a system and its surroundings?
- 3. What are the required characteristics of an isolated system?
- Explain the difference between exothermic and endothermic processes.
- 5. Indicate which of the following processes are exothermic and which are endothermic:
 - a. ice melting
 - b. burning a candle
 - c. cooking an egg.
 - d. the evaporation of sweat
- Calculate the amount of heat needed to increase the temperature of 125 g of water from 22°C to 59°C.
- Calculate the specific heat of copper, given that 204.75 J of energy raises the temperature of 15 g of copper from 35°C to 70°C.
- 432 J of energy is required to raise the temperature of a block of aluminum from 20°C to 60°C. Calculate the mass of aluminum present.
- 300 g of liquid ethanol at 22°C is supplied with 6480 J of heat. What is the final temperature of the ethanol? (The specific heat of ethanol is 2.44 J/g °C).

Answers

- All chemical processes involve the transfer of heat to or from the system.
- The system includes the components involved in a reaction, while the surroundings include everything else.
- 3. An isolated system can exchange neither heat or matter with its surroundings
- Exothermic processes release heat to the surroundings, while endothermic processes require heat from the surroundings.
- a Endothermic
 - b. Exothermic
 - c. Endothermic
 - d. Endothermic
- 6. mass = 125 g ΔT = 37°C specific heat of water = 4.18 J/g °C ΔH = 125g × $\frac{4.18J}{g^{\circ}C}$ × 37°C = 19332.5 J or 19.3 kJ

Pearson Chemistry Answer Key Ch 18

Debabrata Banerjea, M K Bharty

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