

## Answer Key

### Ch. 6 & 7 Review for Test - Chemistry Honors

\* This review worksheet is to help you prepare for the Ch. 6 & 7 Test. It should not be used as your only source of study materials. Past notes, homework assignments, OK-12 Reviews, and warm-ups should be used for a comprehensive review.

*This worksheet will not be submitted. Answer key on Schoology.*

1. All of the following are clues that a chemical reaction has taken place *except*
- a color change
  - a solid forms
  - the reactant is smaller
  - a flame occurs
  - bubbles form

When the following equations are balanced using the smallest possible integers, what is the number in front of the underlined substance in each case?

2.  $\text{Pb}(\text{NO}_3)_2 + \text{K}_2\text{CO}_3 \rightarrow \text{PbCO}_3 + \underline{\text{KNO}_3}$
- A. 5      B. 2      C. 4      D. 3      E. 1

When the following equations are balanced using the smallest possible integers, what is the number in front of the underlined substance in each case?

3.  $\text{KClO}_3 \rightarrow \text{KCl} + \underline{\text{KClO}_2}$
- A. 3      B. 1      C. 2      D. 4      E. 5

4. A substance that, when dissolved in water, produces a solution that conducts electric current very efficiently is called
- an electrical solute
  - a strong ion
  - a strong electrolyte
  - a weak electrolyte
  - None of the above
5. The most common factors that cause chemical reactions to occur are *all* of the following *except*
- formation of a solid
  - formation of a gas
  - formation of water
  - a decrease in temperature
  - transfer of electrons

Identify the solid product that forms when the following aqueous solutions are mixed:

6.  $\text{KNO}_3(\text{aq}) + \text{BaCl}_2(\text{aq}) \rightarrow$
- $\text{KCl}_2$
  - No solid forms.
  - $\text{Ba}(\text{NO}_3)_2$
  - $\text{BaNO}_3$
  - $\text{KCl}$

# Pearson Chemistry Answer Key Ch 6 Test

**Brendan G. Carr**



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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h<sub>2</sub>o

*worked example balancing a redox equation in basic solution* - Jun 23 2022

web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh ions or the oh h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between clo and cr oh<sub>4</sub> in basic solution created by jay

**balancing redox reactions in basic solution problems 1 10** - Aug 06 2023

web bonus problem here is a balanced net ionic equation  $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$  the above equation takes place in two stages first bromine reacts with  $\text{OH}^-$  ions to form bromide ions and  $\text{BrO}_3^-$  ions second the  $\text{BrO}_3^-$  ions react to form bromide ions and bromate  $\text{BrO}_3^-$  ions

### **13b balancing redox equations worksheet chemistry** - Apr 02 2023

web mar 13 2023 balancing redox equations by the ion electron method there are two principal methods for balancing redox equations oxidation state method ion electron method the latter is easier to use with redox reactions in aqueous solution and if necessary can be adapted to many situations that are not in aqueous solution

### **balancing redox equations article khan academy** - Mar 01 2023

web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

### balancing redox reactions in acidic solution problems 1 10 - Sep 07 2023

web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

### **worked example balancing a redox equation in acidic solution** - Mar 21 2022

web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add  $\text{H}^+$  ions or the  $\text{H}_2\text{O}$  pair to fully balance the equation in this video we ll walk through this process for the reaction between dichromate  $\text{Cr}_2\text{O}_7^{2-}$  and chloride  $\text{Cl}^-$  ions in acidic solution created by jay

### *worksheet 5 balancing redox reactions in acid and basic* - Aug 26 2022

web balance each half reaction in basic solution 4 5 6 7  $\text{Cr}_2\text{O}_7^{2-} + 2\text{NO}_2 + \text{SO}_4^{2-} + \text{MnO}_2 \rightarrow \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method 8  $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 9\text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 \rightarrow \text{Cr}^{3+} + \text{TeO}_4 + 11\text{PbO}_2 + \text{IO}_2 + \text{Pb}_2 + 12\text{AsNO}_3 + \text{IO}_3 + \text{ReIO}_3 + \text{H}_2\text{AsO}_4$

### **balancing redox equations practice exercises enig periodic** - Oct 08 2023

web practice exercises balance the following equations of redox reactions assign oxidation numbers to all elements in the reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

### balancing redox reactions examples chemistry libretexts - Jul 05 2023

web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding  $\text{H}_2\text{O}$   $\text{H}^+$  and  $\text{e}^-$



in this order balance the atoms in the equation apart from o and h

**redox reactions questions practice khan academy** - Apr 21 2022

web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$  choose 1 answer  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  no 2 and  $\text{C}_6\text{H}_8\text{O}_6$  b  $\text{NO}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$  h and  $\text{NO}_2$  c h and