

Chapter 7 Chemical Reactions

Section 7.1 Describing Reactions

(pages 192–198)

This section discusses the use of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry.

Reading Strategy (page 192)

Monitoring Your Understanding Preview the Key Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn. After reading, state what you learned about each item you listed. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Expect to Learn	What I Learned
Answers may vary. Possible answers: How to balance chemical equations	Answers may vary. Possible answers: An unbalanced equation can be balanced by changing the coefficients.
How to convert from mass to moles	The mass of a substance can be converted to moles by using the molar mass as a conversion factor.

Chemical Equations (pages 192–193)

- Is the following sentence true or false? The new substances formed as a result of a chemical reaction are called products. true
- Circle the letter of each sentence that is a correct interpretation of the chemical equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$.
 - Carbon and oxygen react and form carbon monoxide.
 - ☒ Carbon and oxygen react and form carbon dioxide.
 - Carbon dioxide yields carbon and oxygen.
 - ☒ The reaction of carbon and oxygen yields carbon dioxide.
- Is the following sentence true or false? The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction. true
- Circle the letter of the correct answer. According to the equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$, how many carbon atoms react with 14 molecules of oxygen to form 14 molecules of carbon dioxide?
 - 1
 - 7
 - ☒ 14
 - 28
- In the reaction represented by the equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$, the mass of carbon dioxide produced equals the total mass of carbon and oxygen that reacted.

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R Sanford



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1. Understanding the eBook Pearson Chemistry Answer Key Ch 7 Vocabulary
 - The Rise of Digital Reading Pearson Chemistry Answer Key Ch 7 Vocabulary
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Answer Key Ch 7 Vocabulary
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Answer Key Ch 7 Vocabulary
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Answer Key Ch 7 Vocabulary
 - Personalized Recommendations
 - Pearson Chemistry Answer Key Ch 7 Vocabulary User Reviews and Ratings
 - Pearson Chemistry Answer Key Ch 7 Vocabulary and Bestseller Lists

5. Accessing Pearson Chemistry Answer Key Ch 7 Vocabulary Free and Paid eBooks
 - Pearson Chemistry Answer Key Ch 7 Vocabulary Public Domain eBooks
 - Pearson Chemistry Answer Key Ch 7 Vocabulary eBook Subscription Services
 - Pearson Chemistry Answer Key Ch 7 Vocabulary Budget-Friendly Options
6. Navigating Pearson Chemistry Answer Key Ch 7 Vocabulary eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Chemistry Answer Key Ch 7 Vocabulary Compatibility with Devices
 - Pearson Chemistry Answer Key Ch 7 Vocabulary Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry Answer Key Ch 7 Vocabulary
 - Highlighting and Note-Taking Pearson Chemistry Answer Key Ch 7 Vocabulary
 - Interactive Elements Pearson Chemistry Answer Key Ch 7 Vocabulary
8. Staying Engaged with Pearson Chemistry Answer Key Ch 7 Vocabulary
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Answer Key Ch 7 Vocabulary
9. Balancing eBooks and Physical Books Pearson Chemistry Answer Key Ch 7 Vocabulary
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Answer Key Ch 7 Vocabulary
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chemistry Answer Key Ch 7 Vocabulary
 - Setting Reading Goals Pearson Chemistry Answer Key Ch 7 Vocabulary
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chemistry Answer Key Ch 7 Vocabulary
 - Fact-Checking eBook Content of Pearson Chemistry Answer Key Ch 7 Vocabulary
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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