

Answers

- 55 $M = 133 \text{ g mol}^{-1}$ so gas is Xe
- 56 90.4 g mol^{-1}
- 57 Helium
- 58 311 dm^3
- 59 empirical formula and molecular formula = SO_2
- 60 At higher altitude the external pressure is less. As the air in the tyre expands on heating (due to friction with the road surface), the internal pressure increases.
- 61 (a) Particles are in constant random motion and collide with each other and with the walls of the container in perfectly elastic collisions. The kinetic energy of the particles increases with temperature. There are no inter-particle forces and the volume of the particles is negligible relative to the volume of the gas.
- (b) At low temperature, the particles have lower kinetic energy, which favours the formation of inter-particle forces and reduces gas pressure. $\frac{PV}{nRT} < 1$
- 62 NH_3 shows greater deviation than CH_4 due to stronger intermolecular attractions, especially at low temperature.
- 63 B
- 64 2.81 g
- 65 4.93 g
- 66 0.0100 mol
- 67 $0.400 \text{ mol dm}^{-3}$
- 68 3.1 cm^3
- 69 $0.178 \text{ mol dm}^{-3}$
- 70 $0.0220 \text{ mol dm}^{-3}$, $0.0602\% \text{ HCl}$
- 71 $0.106 \text{ mol dm}^{-3} \text{ Na}_2\text{SO}_4$ and $0.115 \text{ mol dm}^{-3} \text{ Pb}(\text{NO}_3)_2$; assume no side reactions, all PbSO_4 precipitates
- 72 1217 tonnes

4

- 73 $52\% \text{ NH}_3$ by mass; assuming no side reactions occur and gases behave as ideal gases
- 74 $3.20 \times 10^2 \text{ kg}$
- 75 $0.225 \text{ mol dm}^{-3}$ (or round to 0.23 mol dm^{-3})
- 76 $[\text{Na}_2\text{CO}_3] = \frac{VP}{x} \text{ mol dm}^{-3}$
 $[\text{NaHCO}_3] = \frac{V(P - P)}{x} \text{ mol dm}^{-3}$
- 77 100 mol O_2 ; 3.2 kg

Challenge yourself

- In cold climates, temperature may approach or go below the boiling point of butane so the butane stays liquid even when it is released from the pressure it is under when stored in its canister. This makes it ineffective as a fuel.
- $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$, $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$, $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$
- N = 16%, P = 22%, K = 17%
- Many reactions with 'useless' by-products could have high stoichiometric yield under optimum conditions, but low atom economy, for example, methanoic acid production:

$$2\text{NaCOOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{HCOOH} + \text{Na}_2\text{SO}_4$$
 For 100% conversion with stoichiometric reactants, the yield = 100%.

$$\text{atom economy} = \frac{2 \times 46.03}{(2 \times 68.01) + 98.08} \times 100\% = 39\%$$
- $$2\text{NaN}_3(\text{s}) \rightarrow 2\text{Na}(\text{s}) + 3\text{N}_2(\text{g})$$

$$10\text{Na}(\text{s}) + 2\text{KNO}_3(\text{s}) \rightarrow \text{K}_2\text{O}(\text{s}) + 5\text{Na}_2\text{O}(\text{s}) + \text{N}_2(\text{g})$$

$$\text{K}_2\text{O}(\text{s}) + \text{Na}_2\text{O}(\text{s}) + \text{SiO}_2(\text{s}) \rightarrow \text{Na}_2\text{K}_2\text{SiO}_5 \text{ (alkaline silicate glass)}$$
- As NaOH dissolves, the separated Na^+ and OH^- ions become hydrated, i.e. they are surrounded by H_2O molecules. This involves breaking the hydrogen bonds between the H_2O molecules in pure water and allows closer packing, which reduces the volume.

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Kathleen Armour



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1. Understanding the eBook Pearson Chemistry Chapters 10 Assessment Answers
 - The Rise of Digital Reading Pearson Chemistry Chapters 10 Assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Chapters 10 Assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Chapters 10 Assessment Answers
 - Personalized Recommendations
 - Pearson Chemistry Chapters 10 Assessment Answers User Reviews and Ratings
 - Pearson Chemistry Chapters 10 Assessment Answers and Bestseller Lists
5. Accessing Pearson Chemistry Chapters 10 Assessment Answers Free and Paid eBooks
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 - Pearson Chemistry Chapters 10 Assessment Answers Compatibility with Devices
 - Pearson Chemistry Chapters 10 Assessment Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry Chapters 10 Assessment Answers
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 - Interactive Elements Pearson Chemistry Chapters 10 Assessment Answers
- 8. Staying Engaged with Pearson Chemistry Chapters 10 Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Chapters 10 Assessment Answers
- 9. Balancing eBooks and Physical Books Pearson Chemistry Chapters 10 Assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Chapters 10 Assessment Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chemistry Chapters 10 Assessment Answers
 - Setting Reading Goals Pearson Chemistry Chapters 10 Assessment Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chemistry Chapters 10 Assessment Answers
 - Fact-Checking eBook Content of Pearson Chemistry Chapters 10 Assessment Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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