Reviewing Content

- The solvent is the substance in which the solute is dissolved.
- Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
- 44. solubility: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.

saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. unsaturated solution: a solution that contains less solute than a saturated solution at a given temperature and pressure. miscible: describes liquids that dissolve in each other. immiscible: describes liquids that are insoluble in each other.

- 45. Particles of solute crystallize.
- No: if there were undissolved solute, the excess solute would come out of a supersaturated solution.
- 47. 5.55 × 102 g AgNO₃
- 48. Solubility increases with pressure.
- 49. a. 1.6×10⁻² g/L b. 4.7×10⁻² g/L
- Dilute and concentrated are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
- Molarity is the number of moles of solute dissolved in one liter of solution.
 - a. 1.3M KCI
 - b. 3.3 × 10⁻¹M MgCl₂
- 52. 2.00×10^{1} mL.
- a. 5.0×10⁻¹ mol NaCl, 29 g NaCl
 b. 1.0 mol KNO₃, 1.0×10⁻² g KNO₃
 c. 2.5×10² mol CaCl₂, 2.8 g CaCl₂
- 54. a. 2.3 × 10¹ g NaCl b. 2.0 g MgCl.
- 55. a. 16% (v/v) ethanol
 - b. 63.6% (v/v) isopropyl alcohol

- Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation. freezing-point depression, and vaporpressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution: relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
- 57. a. seawater
 - b. 1.5M KNO₃
 - e. 0.100M MgCl₂
- The effective molality of the Ca(NO₃)₂ solution is 3m. The effective molality of the NaNO₃ solution is 2m.
- 59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
- The salt lowers the freezing point of the ice-water cooling mixture.
- 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
- 62. Add 27.0 g H2O to 32.0 g CH2OH.
- 63. a. 100.26°C
 - b. 101.54°C
- 64. a. -4.46°C
 - b. -2.2°C
- 65. a. -1.1°C
 - b. -0.74°C
 - e. −1.5°C

Understanding Concepts

 a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

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Antony C. Wilbraham

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 - The Rise of Digital Reading Pearson Chemistry Chapters 16 Assessment Answers
 - Advantages of eBooks Over Traditional Books
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 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Chapters 16 Assessment Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Pearson Chemistry Chapters 16 Assessment Answers
 - Personalized Recommendations
 - Pearson Chemistry Chapters 16 Assessment Answers User Reviews and Ratings
 - Pearson Chemistry Chapters 16 Assessment Answers and Bestseller Lists

- 5. Accessing Pearson Chemistry Chapters 16 Assessment Answers Free and Paid eBooks
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- 8. Staying Engaged with Pearson Chemistry Chapters 16 Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Chapters 16 Assessment Answers
- 9. Balancing eBooks and Physical Books Pearson Chemistry Chapters 16 Assessment Answers
 - Benefits of a Digital Library
 - \circ Creating a Diverse Reading Collection Pearson Chemistry Chapters 16 Assessment Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chemistry Chapters 16 Assessment Answers
 - Setting Reading Goals Pearson Chemistry Chapters 16 Assessment Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chemistry Chapters 16 Assessment Answers
 - Fact-Checking eBook Content of Pearson Chemistry Chapters 16 Assessment Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

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- Exploring Educational eBooks
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 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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