

6

THE PERIODIC TABLE

Chapter Quiz

Fill in the word(s) that will make each statement true.

- | | |
|--|--------------|
| 1. In the modern periodic table, when elements are arranged according to their atomic <u>1</u> , there is a periodic repetition of properties. | 1. _____ 6.1 |
| 2. There are <u>2</u> periods in the periodic table. | 2. _____ 6.1 |
| 3. The elements in any <u>3</u> in the periodic table have similar physical and chemical properties. | 3. _____ 6.1 |
| 4. Oxygen and sulfur, Group 6A elements, have <u>4</u> electrons in their highest occupied energy level. | 4. _____ 6.2 |
| 5. For the inner transition elements, electrons are added to an <i>f</i> sublevel with a principal energy level that is <u>5</u> than the period number. | 5. _____ 6.2 |

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- | | |
|---|-----|
| _____ 6. The atomic radius of an element in period 3 is larger than the atomic radius of an element in period 2. | 6.3 |
| _____ 7. For Group 3A elements, there is a relatively small increase in ionization energy between the second and third ionization energies. | 6.3 |
| _____ 8. Anions are smaller than the neutral atoms from which they are formed. | 6.3 |
| _____ 9. Atoms with low electronegativity values tend to form positive ions. | 6.3 |
| _____ 10. As a group, alkali metals have the highest electronegativities. | 6.3 |

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L. G. Wade



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