

Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility**: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.
saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution**: a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible**: describes liquids that dissolve in each other. **immiscible**: describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47. 5.55×10^2 g AgNO_3
48. Solubility increases with pressure.
49. a. 1.6×10^{-2} g/L
b. 4.7×10^{-2} g/L
50. *Dilute* and *concentrated* are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.
a. 1.3M KCl
b. 3.3×10^{-1} M MgCl_2
52. 2.00×10^3 mL
53. a. 5.0×10^{-3} mol NaCl, 29 g NaCl
b. 1.0 mol KNO_3 , 1.0×10^{-2} g KNO_3
c. 2.5×10^3 mol CaCl_2 , 2.8 g CaCl_2
54. a. 2.3×10^1 g NaCl
b. 2.0 g MgCl_2
55. a. 16% (v/v) ethanol
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water
b. 1.5M KNO_3
c. 0.100M MgCl_2
58. The effective molality of the $\text{Ca}(\text{NO}_3)_2$ solution is 3m. The effective molality of the NaNO_3 solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g H_2O to 32.0 g CH_3OH .
63. a. 100.26°C
b. 101.54°C
64. a. -4.46°C
b. -2.2°C
65. a. -1.1°C
b. -0.74°C
c. -1.5°C

Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

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An Introduction to Medicinal Chemistry Graham L. Patrick, 2013-01-10 This volume provides an introduction to medicinal chemistry It covers basic principles and background and describes the general tactics and strategies involved in developing an effective drug

Chemistry Pearson Education, Lemay, 2005-05 This lecture notebook contains the art from the text with note taking sections to obviate the need for students to spend time redrawing figures in lecture and instead lets them concentrate on taking notes

Survey of Industrial Chemistry Philip J. Chenier, 2012-12-06 Survey of Industrial Chemistry arose from a need for a basic text dealing with industrial chemistry for use in a one semester three credit senior level course taught at the University of Wisconsin Eau Claire This edition covers all important areas of the chemical industry yet it is reasonable that it can be covered in 40 hours of lecture Also an excellent resource and reference for persons working in the chemical and related industries it has sections on all important technologies used by these industries a one step source to answer most questions on practical applied chemistry Young scientists and engineers just entering the workforce will find it especially useful as a readily available handbook to prepare them for a type of chemistry quite different than they have seen

in their traditional coursework whether graduate or undergraduate

Organic Chemistry Pierre Vogel, Kendall N. Houk, 2019-10-07 Provides the background tools and models required to understand organic synthesis and plan chemical reactions more efficiently Knowledge of physical chemistry is essential for achieving successful chemical reactions in organic chemistry Chemists must be competent in a range of areas to understand organic synthesis Organic Chemistry provides the methods models and tools necessary to fully comprehend organic reactions Written by two internationally recognized experts in the field this much needed textbook fills a gap in current literature on physical organic chemistry Rigorous yet straightforward chapters first examine chemical equilibria thermodynamics reaction rates and mechanisms and molecular orbital theory providing readers with a strong foundation in physical organic chemistry Subsequent chapters demonstrate various reactions involving organic organometallic and biochemical reactants and catalysts Throughout the text numerous questions and exercises over 800 in total help readers strengthen their comprehension of the subject and highlight key points of learning The companion Organic Chemistry Workbook contains complete references and answers to every question in this text A much needed resource for students and working chemists alike this text Presents models that establish if a reaction is possible estimate how long it will take and determine its properties Describes reactions with broad practical value in synthesis and biology such as C C coupling reactions pericyclic reactions and catalytic reactions Enables readers to plan chemical reactions more efficiently Features clear illustrations figures and tables With a Foreword by Nobel Prize Laureate Robert H Grubbs Organic Chemistry Theory Reactivity and Mechanisms in Modern Synthesis is an ideal textbook for students and instructors of chemistry and a valuable work of reference for organic chemists physical chemists and chemical engineers

Microchemical Engineering in Practice Thomas Dietrich, 2011-11-18 Microchemical Engineering in Practice provides the information chemists and engineers need to evaluate the use of microreactors covering the technical operational and economic considerations for various applications It explains the systems needed to use microreactors in production and presents examples of microreactor use in different chemistries including larger scale production processes There are guidelines on calculating the costs and the risks of production using continuous flow microreactors Complete with case studies this is an essential guide for chemists and engineers interested in investigating the advantages of chemical microreactors

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