Reviewing Content

- The solvent is the substance in which the solute is dissolved.
- Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
- 44. solubility: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.

saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. unsaturated solution: a solution that contains less solute than a saturated solution at a given temperature and pressure. miscible: describes liquids that dissolve in each other. immiscible: describes liquids that are insoluble in each other.

- 45. Particles of solute crystallize.
- No: if there were undissolved solute, the excess solute would come out of a supersaturated solution.
- 47. 5.55 × 102 g AgNO₃
- 48. Solubility increases with pressure.
- 49. a. 1.6×10⁻² g/L b. 4.7×10⁻² g/L
- Dilute and concentrated are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
- Molarity is the number of moles of solute dissolved in one liter of solution.
 - a. 1.3M KCI
 - b. 3.3 × 10⁻¹M MgCl₂
- 52. 2.00×10^{1} mL.
- a. 5.0×10⁻¹ mol NaCl, 29 g NaCl
 b. 1.0 mol KNO₃, 1.0×10⁻² g KNO₃
 c. 2.5×10² mol CaCl₂, 2.8 g CaCl₂
- 54. a. 2.3 × 10¹ g NaCl b. 2.0 g MgCl.
- 55. a. 16% (v/v) ethanol
 - b. 63.6% (v/v) isopropyl alcohol

- Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation. freezing-point depression, and vaporpressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution: relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
- 57. a. seawater
 - b. 1.5M KNO₃
 - e. 0.100M MgCl₂
- The effective molality of the Ca(NO₃)₂ solution is 3m. The effective molality of the NaNO₃ solution is 2m.
- 59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
- The salt lowers the freezing point of the ice-water cooling mixture.
- 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
- 62. Add 27.0 g H2O to 32.0 g CH2OH.
- 63. a. 100.26°C
 - b. 101.54°C
- 64. a. -4.46°C
 - b. -2.2°C
- 65. a. -1.1°C
 - b. -0.74°C
 - e. −1.5°C

Understanding Concepts

 a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

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David Kirk

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Teaching Innovation in University Education: Case Studies and Main Practices Saura, Jose Ramon, 2022-06-17 In the last decade the development of new technologies has made innovation a fundamental pillar of education Teaching innovation includes the evolution of both teaching and learning models to drive improvements in educational methodologies Teaching innovation is a pioneer in the understanding and comprehension of the different teaching methodologies and models developed in the academic area Teaching innovation is a process that seeks validation in the academic and teaching communities at universities in order to promote the improvement and its practices and uses in the future characterized by digital development and data based methods Teaching Innovation in University Education Case Studies and Main Practices features the major practices and case studies of teaching innovation developed in recent years at universities It is a source on study cases focused on teaching innovation methodologies as well as on the identification of new technologies that will help the development of initiatives and practices focused on teaching innovation at higher education institutions Covering topics such as didactic strategics service learning and technology based gamification this premier reference source is an indispensable resource for pre service teachers lecturers students faculty administrators libraries entrepreneurs researchers Chemistry: The Key to our Sustainable Future Minu Gupta Bhowon, Sabina and academicians Jhaumeer-Laulloo, Henri Li Kam Wah, Ponnadurai Ramasami, 2013-11-08 Chemistry The Key to our Sustainable Future is a collection of selected contributed papers by participants of the International Conference on Pure and Applied Chemistry ICPAC 2012 on the theme of Chemistry The Key for our Future held in Mauritius in July 2012 In light of the significant contribution of chemistry to benefit of mankind this book is a collection of recent results generated from research in chemistry and interdisciplinary areas It covers topics ranging from nanotechnology natural product chemistry to analytical and environmental chemistry Chemistry The Key to our Sustainable Future is written for graduates postgraduates researchers in industry and academia who have an interest in the fields ranging from fundamental to applied chemistry

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