

## Pearsons Learning : Chapter 14 : Solutions

1. What is the volume/volume percent of a solution prepared from 12.3 mL ethanol in 150.0 mL water? Assume volumes are additive.

Answer : percent concentration =  $\frac{\text{amount of solute}}{\text{amount of solute} + \text{amount of solvent}} \times 100$

$$\text{percent concentration (ethanol)} = \frac{12.3 \text{ mL ethanol}}{150.0 \text{ mL} + 12.3 \text{ mL}} \times 100 = 7.58\%$$
$$100\% - 92.4\% = \boxed{7.58\% \text{ (v/v) ethanol}}$$

2. Which expression is equivalent to 5.45% (v/v) ethanol?

Answer : percent concentration =  $\frac{\text{amount of solute}}{\text{amount of solute} + \text{amount of solvent}} \times 100$

$$5.45\% \text{ (v/v) ethanol} \div 100 = 0.0545 \text{ mL}$$
$$0.0545 \text{ mL} = \boxed{\frac{5.45 \text{ mL ethanol}}{100 \text{ mL solution}}}$$

3. What mass of NaCl is present in 50.0 g of 3.75% (m/m) NaCl solution?

Answer :  $3.75\% \text{ (m/m) NaCl} \div 100 = 0.0375 \text{ g}$

$$\text{mass percent concentration} = \frac{\text{mass of solute}}{\text{mass of solute} + \text{mass of solvent}} \times 100$$
$$96.376\% = \frac{50.0 \text{ g}}{50.0 \text{ g} + 1.88 \text{ g}}$$
$$100\% - 96.376\% = 3.624\% \text{ (m/m) NaCl}$$
$$\text{Mass of NaCl} = \boxed{1.88 \text{ g}}$$

4. What is the osmotic pressure of a 0.850 M solution of glucose in water at 35 °C?

# Pearson Chapter 14 Mixture And Solutions Worksheet

**David B. Troy, Paul Beringer**



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## **Table of Contents Pearson Chapter 14 Mixture And Solutions Worksheet**

1. Understanding the eBook Pearson Chapter 14 Mixture And Solutions Worksheet
  - The Rise of Digital Reading Pearson Chapter 14 Mixture And Solutions Worksheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chapter 14 Mixture And Solutions Worksheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson Chapter 14 Mixture And Solutions Worksheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chapter 14 Mixture And Solutions Worksheet
  - Personalized Recommendations
  - Pearson Chapter 14 Mixture And Solutions Worksheet User Reviews and Ratings
  - Pearson Chapter 14 Mixture And Solutions Worksheet and Bestseller Lists
5. Accessing Pearson Chapter 14 Mixture And Solutions Worksheet Free and Paid eBooks
  - Pearson Chapter 14 Mixture And Solutions Worksheet Public Domain eBooks
  - Pearson Chapter 14 Mixture And Solutions Worksheet eBook Subscription Services
  - Pearson Chapter 14 Mixture And Solutions Worksheet Budget-Friendly Options

6. Navigating Pearson Chapter 14 Mixture And Solutions Worksheet eBook Formats
  - ePub, PDF, MOBI, and More
  - Pearson Chapter 14 Mixture And Solutions Worksheet Compatibility with Devices
  - Pearson Chapter 14 Mixture And Solutions Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pearson Chapter 14 Mixture And Solutions Worksheet
  - Highlighting and Note-Taking Pearson Chapter 14 Mixture And Solutions Worksheet
  - Interactive Elements Pearson Chapter 14 Mixture And Solutions Worksheet
8. Staying Engaged with Pearson Chapter 14 Mixture And Solutions Worksheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pearson Chapter 14 Mixture And Solutions Worksheet
9. Balancing eBooks and Physical Books Pearson Chapter 14 Mixture And Solutions Worksheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pearson Chapter 14 Mixture And Solutions Worksheet
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chapter 14 Mixture And Solutions Worksheet
  - Setting Reading Goals Pearson Chapter 14 Mixture And Solutions Worksheet
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chapter 14 Mixture And Solutions Worksheet
  - Fact-Checking eBook Content of Pearson Chapter 14 Mixture And Solutions Worksheet
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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