

**6****THE PERIODIC TABLE**

## Chapter 6 Review Package - Part 1

**A. Matching**

Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line.

**Column A**

- \_\_\_\_\_ 1. metal
- \_\_\_\_\_ 2. halogen
- \_\_\_\_\_ 3. inner transition metal
- \_\_\_\_\_ 4. representative element
- \_\_\_\_\_ 5. ~~ionization energy~~
- \_\_\_\_\_ 6. alkaline earth metal
- \_\_\_\_\_ 7. cation
- \_\_\_\_\_ 8. electronegativity
- \_\_\_\_\_ 9. transition metal
- \_\_\_\_\_ 10. noble gas

**Column B**

- a. positively charged ion
- b. an element whose highest occupied *s* or *p* sublevels are partially filled
- c. ~~the energy needed to remove an electron from an atom in the gaseous state~~
- d. good conductor of heat and electric current
- e. an element whose highest occupied *s* sublevel and a nearby *d* sublevel contain electrons
- f. any nonmetal in Group 7A
- g. ~~the tendency of an atom to attract electrons when the atom is in a compound~~
- h. an element in which the highest occupied *s* and *p* sublevels are filled
- i. an element in Group 2A
- j. an element whose highest occupied *s* sublevel and a nearby *f* sublevel contain electrons

**B. Multiple Choice**

Choose the best answer and write its letter on the line.

- \_\_\_\_\_ 11. In the periodic table, there is a periodic pattern in the physical and chemical properties of elements when they are arranged in order of
  - a. increasing atomic mass.
  - b. increasing electronegativity.
  - c. increasing atomic radius.
  - d. increasing atomic number.

# Pearson Chemistry Chapter 6 The Periodic Table Test

**Sabine Zange**



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