

CK-12 Chemistry - Basic Answer Key

Chapter 7: Chemical Nomenclature

7.1 Ionic Compounds

Lesson Review Questions

Questions

- What is the purpose of an empirical formula?
- Define binary and ternary ionic compounds.
- What is the difference between a monatomic and a polyatomic ion?
- For each of the following ionic compounds, what ions are present and in what ratio?
 - MgBr_2
 - Li_2CO_3
 - $\text{Fe}_2(\text{SO}_4)_3$
- Predict the compound that forms when gallium combines with oxygen. What would the name of this compound be?
- Give the formula for each of the following ionic compounds:

(a) ammonium nitrate	(d) vanadium(III) oxide
(b) cobalt(II) sulfate	(e) barium oxide
(c) nickel(II) cyanide	(f) calcium hypochlorite
- Name the following ionic compounds:
 - MgBr_2
 - (d) KMnO_4
 - Li_2CO_3
 - $(\text{NH}_4)_2\text{S}$
 - KHSO_3
 - CuCl
 - CuCl_2
- Write the correct formulas for the following ionic compounds:
 - Barium chloride
 - Chromium(III) oxide
 - Potassium sulfate
 - Zinc phosphate

Answer:

- The empirical formula is the most straightforward way to describe the structure of a molecule, using a chemical formula that gives the lowest whole number ratio between the ions involved.

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Host-Guest Chemistry Brian D. Wagner, 2020-10-26 This textbook addresses the chemical and physicochemical principles of supramolecular host guest chemistry in solution It covers the thermodynamics and dynamics of inclusion and highlights several types of organic hosts Various applications of host guest chemistry in analytical and environmental chemistry as well as pharmaceutical and chemical industry demonstrate the versatile usability of molecular cages

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