

- 13 These calculations have used $L = 6.02 \times 10^{23}$
 (a) 7.22×10^{22} (b) 3.01×10^{23}
 (c) 1.20×10^{23}

14 0.53 mol H

15 0.250 mol

- 16 (a) 262.87 g mol⁻¹ (b) 176.14 g mol⁻¹
 (c) 164.10 g mol⁻¹ (d) 248.22 g mol⁻¹

17 189.1 g

18 1.5 mol

19 0.0074 mol Cl⁻

20 1.83×10^{23} C atoms

21 171 g (integer value because no calculator)

22 10.0 g H₂O

23 2.0 mol N₂ > 3.0 mol NH₃ > 25.0 mol H₂ >
 1.0 mol N₂H₄

- 24 (a) CH (b) CH₂O
 (c) C₂H₄O₂ (d) C₂H₆
 (e) C₂H₄ (f) CH₃O

25 Na₂S₂O₅

26 CoSO₄·7H₂O

27 C₁₀H₁₆N

28 NH₃

29 6.94 Li

30 CdS

31 empirical formula CH; molecular formula C₇H₈

32 empirical formula H₂PO₃; molecular formula H₄P₂O₆

33 C₁₀H₁₆N₂P₂O₁₀ for both empirical and molecular formulas

34 C₂H₃O

35 Let y = mass of chalk in grams.

$$\text{moles of chalk used} = \frac{\text{mass used}}{M_r(\text{CaCO}_3)} \times$$

$$y \frac{1 \text{ g}}{100.09 \text{ g mol}^{-1}}$$

This is the same as the number of moles of carbon atoms used.

$$\text{Therefore the number of carbon atoms used} = \text{moles of chalk} \times (6.02 \times 10^{23} \text{ mol}^{-1}) = \frac{6.02 \times 10^{23} y}{100.09}$$

- 36 (a) 2.50 mol (b) 5.63 mol
 (c) 665.5 g

37 (a) $2\text{C}_2\text{H}_6 + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$
 (b) 1.59 g

38 4.355 kg

- 39 (a) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
 (b) 92.8%

(c) CaCO₃ is the only source of CO₂; all the CaCO₃ undergoes complete decomposition; all CO₂ released is captured; heating does not cause any change in the mass of the other minerals present.

- 40 (a) 65.2 g (b) 1.3 g H₂

41 5.23 g C₂H₄Cl₂

42 253.2 g theoretical CaSO₄; 77.9%

43 3.16 g ester

44 107 g of C₂H₆ needed

- 45 (a) 2.40 mol (b) 0.0110 mol
 (c) 44 mol

- 46 (a) 35.65 dm³ (b) 5.7 dm³

47 0.652 dm³

48 0.138 mol Br₂ and 0.156 mol Cl₂, so more molecules of Cl₂

49 0.113 dm³

50 0.28 dm³

51 90 kPa

52 16 °C

53 3.0 dm³

54 2.8 dm³

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JoAnn Zerwekh, Jo Carol Claborn



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Chemistry insights 'O' level Rex M. Heyworth, 2007 **Clinical Chemistry - E-Book** Donna Larson, 2015-12-17 Gain a clear understanding of pathophysiology and lab testing Clinical Chemistry Fundamentals and Laboratory Techniques prepares you for success as a medical lab technician by simplifying complex chemistry concepts and lab essentials including immunoassays molecular diagnostics and quality control A pathophysiologic approach covers diseases that are commonly diagnosed through chemical tests broken down by body system and category such as respiratory gastrointestinal and cardiovascular conditions Written by clinical chemistry educator Donna Larson and a team of expert contributors this full color book is ideal for readers who may have minimal knowledge of chemistry and are learning laboratory science for the first time Full color illustrations and design simplify complex concepts and make learning easier by highlighting important material Case studies help you apply information to real life scenarios Pathophysiology and Analytes section includes information related to diseases or conditions such as a biochemistry review disease mechanisms clinical correlation and laboratory analytes and assays Evolve companion website includes case studies and animations that reinforce what you ve learned from the book Laboratory Principles section covers safety quality assurance and other fundamentals of laboratory techniques Review questions at the end of each chapter are tied to the learning objectives helping you review and retain the material Critical thinking questions and discussion questions help you think about and apply key points and concepts Other Aspects of Clinical Chemistry section covers therapeutic drug monitoring toxicology transplantation and emergency preparedness Learning objectives in each chapter help you to remember key points or to analyze and synthesize concepts in clinical chemistry A list of key words Is provided at the beginning of each chapter and these are also bolded in the text Chapter summaries consist of bulleted lists and tables highlighting the most important points of each chapter A glossary at the back of the book provides a quick reference to definitions of all clinical chemistry terms **X-Kit Physiology** , 2006

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