

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, $\text{KClO}_3(\text{s})$. Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{23} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(II) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. Write a balanced equation first.

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of dinitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



Pearson Chemistry Workbook Chapter 12 Stoichiometry

James Joseph Gallagher



Pearson Chemistry Workbook Chapter 12 Stoichiometry:

Teaching Science for Understanding James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities Students Study Guide Fredeen, Pearson, 2003-07 Formulation and Stoichiometry Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction **Stoichiometry** Sydney Young, 1918 This introduction was originally prefixed to Dr Findlay's Phase rule which was the first volume of the series issued It belongs properly however to this volume and is therefore included here Note p vii Introduction issued also separately in 1904 **Mole Concepts and Stoichiometry** Dharshi Bopegedera, 2017 Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples *Pearson Chemistry 12 Qld Student Book*, 2019 Chemistry Unit 3 Jack Anderson, Georgia Anderson, Tyler Clement, Aryan Puri, Rabea Zeynali, 2020-12-20 ThrivEd is a student run not for profit organisation that writes textbooks on the philosophy of writing textbooks that truly empower students Our Unit 3 Chemistry textbook is a 350 page textbook that covers Chemical Equilibrium Stoichiometry Acids Bases Volumetric Analysis and Redox The proceeds from every two books sold are used to produce and donate a third textbook for a student who would benefit from the same support The book is structured in five main sections 1 Bridging the Gap To create the connections needed to set off on the right foot in Year 12 the Bridging the Gap sections are designed to provide in depth teachings of the foundational Year 11 content 2 Main Content By extensively working through concepts with no assumptions and teaching with clear explanations worked examples and real world applications we hope to make each chapter an exciting empowering and well connected experience 3 Topic Test and Exam Notes To guide your preparation of

personal notes leading up to assessments sample test and exam notes have been provided to make this process effective and intentional 4 Progressive Repetitive and Advanced Questions Each progressive question is designed to progressively teach each concept and is paired with multiple repetitive questions allowing you to stop and practice anything you like Advanced questions are also there for some extra fun 5 Online Answer Guide With each question comes a three stage answer including a sample response points to note about the question and a WACE inspired marking guide allowing each question to be understood in all its elements These answer guides are free as a downloadable PDFs It is with this structure that the Unit is extensively covered and targeted to provide the greatest possible benefits to students Beginning Chemistry Anne

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1. Understanding the eBook Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - The Rise of Digital Reading Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - Personalized Recommendations
 - Pearson Chemistry Workbook Chapter 12 Stoichiometry User Reviews and Ratings

- Pearson Chemistry Workbook Chapter 12 Stoichiometry and Bestseller Lists
- 5. Accessing Pearson Chemistry Workbook Chapter 12 Stoichiometry Free and Paid eBooks
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 - Pearson Chemistry Workbook Chapter 12 Stoichiometry Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Pearson Chemistry Workbook Chapter 12 Stoichiometry Compatibility with Devices
 - Pearson Chemistry Workbook Chapter 12 Stoichiometry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - Highlighting and Note-Taking Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - Interactive Elements Pearson Chemistry Workbook Chapter 12 Stoichiometry
- 8. Staying Engaged with Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Workbook Chapter 12 Stoichiometry
- 9. Balancing eBooks and Physical Books Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Workbook Chapter 12 Stoichiometry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - Setting Reading Goals Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - Fact-Checking eBook Content of Pearson Chemistry Workbook Chapter 12 Stoichiometry
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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