

6**Chemical Quantities**

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A. Completion

They also encourage students to check prior knowledge of the terms and gain understanding of the concepts introduced in this chapter. Each blank can be completed with a unit, three places, or nothing.

Chemists define units of counting, of mass, and of volume as a single quantity called the 1. Chemists measure the "amount of substance." One way to measure the amount of a substance is to count the number of 2 of that substance. The amount of representative particles for a sample of a substance is called 3. This number is 4. For example, a mole of nitrogen contains $6.02 \times 10^{23} 5 of nitrogen.$

Chemists define gram atomic mass as the weight of grams of an element whose mass is equal to the 6 atomic mass. This is the mass of $6.02 \times 10^{23} of carbon, equal to the atomic weight of carbon as 7 g of carbon atoms.$

To find the amount in moles of a substance represented, we will suppose the 8 of the above-stating equal analogy. When the substance the mole gives the mass, we divide the 9 of the compound. If we divide a compound that is not an appropriate quantity, because the representative particles found in one mole is composed in its decomposition, the appropriate number is called as 10. The term gram formula mass and 11 can be used to refer to the mass of a mole of any substance or compound.

The gram formula mass is a useful factor in computing a relationship between the number of a substance/compound. The relationship that number of moles is a compound/mole compound, for example, we can find gram formula mass as 12. For example, find all the types, find the value for the number the number 13 for the substance.

1. mass2. representative particles3. Avogadro's Number4. 6.02×10^{23} 5. molecules6. atomic mass7. 128. gram atomic masses9. gram molecular mass10. gram formula mass11. Molar mass12. conversion factor13. chemical formula

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